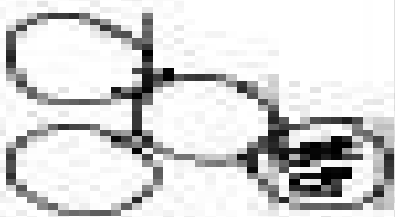


ChemThink: Molecular Shapes

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Understanding that, as they travel, all molecules that have lone pairs have a geometry of electron with lone pairs each atom and that these repulsion will affect the shape of the molecule.

Number of e- surrounding atom	Electron pair arrangement	Bond Angles	Shape	E.g. Lewis dot representation
2		180°	Linear	$H-H$
4		109.5°	Tetrahedral Trigonal Pyramidal	
6		90°	Square	

Answers To Chemthink Molecular Shapes

Guillaume Favre



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