

Name: _____ Per: _____

AP CHEM Gas Laws WorksheetUse the ideal gas law, **PV = nRT** and the universal gas constants:

$$R = 62.36 \frac{\text{L} \cdot \text{torr}}{\text{K} \cdot \text{mol}}$$

or

$$R = 0.08206 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}}$$

$$8.314 \frac{\text{J}}{\text{K} \cdot \text{mol}} = \text{kPa} \cdot \text{L} \cdot \text{mol}^{-1}$$

Combined Gas Laws:
$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

1) Four moles of a gas are at a pressure of 5.6 atm and a volume of 12 liters. What is the temperature?

$$\frac{PV}{nR} = \frac{5.6 \text{ atm} \times 12 \text{ L}}{4 \times 0.0821} = \frac{67.2}{33.256} = 204.73 \text{ K}$$

$$= -68.41^\circ\text{C}$$

2) There is an unknown quantity of gas at a pressure of 1.2 atm, a volume of 31 liters, and a temperature of 87 °C. How many moles of gas do you have?

$$n = \frac{PV}{RT}$$

$$= \frac{1.2 \text{ atm} \times 31 \text{ L}}{360.15 \text{ K} \times 0.0821}$$

$$= \frac{37.2}{29.54} = 1.26 \text{ mol}$$

3) 2.4 moles of gas are held at a temperature of 97 °C in a container with a volume of 45 liters, what is the pressure of the gas?

$$P = \frac{2.4 \times 8.21 \times 370.15}{45}$$

$$= 8.2 \text{ kPa}$$

4) A constant volume of oxygen is heated from 100°C to 185°C. The initial pressure is 4.1 atm. What is the final pressure?

$$P_2 = \frac{P_1 T_2 V_1}{T_1 V_2}$$

$$P_2 = \frac{4.1 \text{ atm} \times 458.15 \text{ K}}{373.15 \text{ K}}$$

$$= 5.0 \text{ kPa}$$

$$= 5.0 \text{ atm}$$

5) A sample of 25L of NH₃ gas at 10°C is heated at constant pressure until it fills a volume of 50L. What is the new temperature in °C?

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

$$T_2 = \frac{282.15 \text{ K} \times 25}{50}$$

$$= 141.575 \text{ K}$$

$$= -131.4^\circ\text{C}$$

Answers To Gas Laws Chem 170

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Answers To Gas Laws Chem 170:

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