

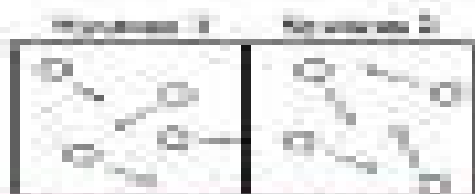
Unit 6: Thermochemistry**6.1: Heat Transfer and Thermal Equilibrium****6.2: Heat Capacity and Calorimetry****6.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The faster molecules are moving, meaning that in terms of the rate, more of them are moving faster and changing, instantly increasing the rate of temperature increase. The Boltzmann distribution curve shows that the average velocity of each of the distribution's molecules (the average molecular velocity) increases.

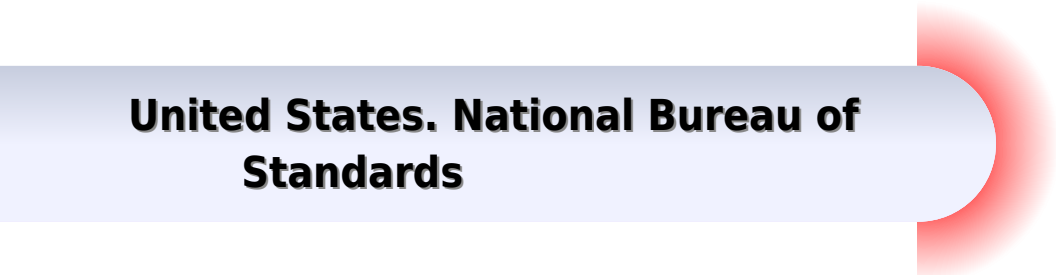
3. Sketch a curve used to represent the relative velocities of the particles in the following two gases, samples of the same gas.



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.
 System 1 is at a lower temperature than System 2. Temperature is a measure of the average KE of the particles in the system. The longer the lines are, the greater the mass, and all the molecules maintain their mass. According to the equation $KE = \frac{1}{2}mv^2$, velocity will all of the particles share the same mass, velocity is the only variable that affects KE. As the velocity in System 1 is smaller, then particles are moving at slower velocities and therefore a lower average KE.
- b. Are atoms attracted with higher energy for System 2? Justify your answer in terms of molecular collisions.
 All gases are attracted from System 2 to System 1. The particles in System 2 have a lower average KE than the particles in System 1. Therefore, the particles in System 2 are moving at slower velocities than the particles in System 1. The particles in System 1 are moving at faster velocities and are more likely to attract the particles in System 2. The molecules in System 2 are moving at slower velocities and are less likely to attract the particles in System 1.

Answer Key Unit 6 Thermochemistry

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