

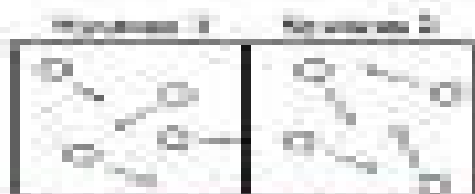
Unit 5: Thermochemistry**5.1: Heat Transfer and Thermal Equilibrium****5.2: Heat Capacity and Calorimetry****5.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving 100 m/s is 100 J . The particles at different temperatures have the same kinetic energy, but the number of particles with that energy changes, resulting in different temperatures. The Boltzmann distribution curve shows that the number of particles at each energy level is proportional to the number of particles with that energy.

3. Sketch a curve to represent the relative distribution of the particles in the following two gas samples at the same gas:



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.
- System 1 is at a higher temperature than System 2. Temperature is a measure of the average kinetic energy of the particles in a system. The particles in System 1 are moving faster than the particles in System 2, so the average kinetic energy of the particles in System 1 is higher than the average kinetic energy of the particles in System 2. Therefore, System 1 is at a higher temperature than System 2.
- b. Are there different total kinetic energy for System 1? Justify your answer in terms of molecular collisions.
- Yes, there is a difference in total kinetic energy between System 1 and System 2. System 1 has a higher temperature than System 2, so the particles in System 1 are moving faster than the particles in System 2. This means that the total kinetic energy of the particles in System 1 is higher than the total kinetic energy of the particles in System 2.

Answer Key Unit 6 Thermochemistry

Adrian Dingle



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