

SCH4U Unit 4 - Quiz – Equilibrium

Name: _____ Period: _____

1. (2)

2. (2)

3. (2)

4. (2)

5. (2)

6. (2) (2) (2) (2)

$$K_c = \frac{[\text{products}]}{[\text{reactants}]} \quad K_p = \frac{(P)^{\text{products}}}{(P)^{\text{reactants}}}$$

$$K_c = \frac{P_{\text{H}_2\text{O}}^2 P_{\text{CO}_2}}{P_{\text{CH}_4} P_{\text{CO}}^2} \quad K_p = \frac{(P_{\text{H}_2\text{O}})^2 (P_{\text{CO}_2})}{(P_{\text{CH}_4}) (P_{\text{CO}})^2}$$

gases: $\Delta n_{\text{gas}} = 2 + 1 - 3 = 0$ $Q_p = \frac{(P_{\text{H}_2\text{O}})^2 (P_{\text{CO}_2})}{(P_{\text{CH}_4}) (P_{\text{CO}})^2} = K_p = 0.26$ $Q_p < K_p$ so $K_p > Q_p$

Question 1:

1. Hydrogen sulfide, H_2S , is a slightly poisonous gas. At 1000°C, an equilibrium mixture contains 0.020 moles of hydrogen, H_2 , 0.010 moles of hydrogen sulfide, with an equilibrium pressure of 0.026 atm. The reaction is:



If the volume of an 8.00 L container at 1000°C where a concentration of H_2 is present at equilibrium:

- What is the concentration of H_2 at equilibrium?
 Answer: $[\text{H}_2] = 0.0025 \text{ mol/L}$
 Answer: $[\text{H}_2\text{S}] = 0.005 \text{ mol/L}$
 $K_c = 0.026$
- Determine concentration of S_2 at equilibrium.
- Answer:
 What other species have chemical formula consisting of 2 atoms conjugate ions?
- Calculate:

	$[\text{H}_2\text{S}]_{\text{eq}}$	$[\text{H}_2]_{\text{eq}}$	$[\text{S}_2]_{\text{eq}}$
Initial concentration (moles/L)			
Change in concentration (moles/L)			
Equilibrium concentration	0.005	0.010	0.0025

1. Not independent equilibrium.

Answers For Sch4u Unit 4

Léo Pomar



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