

SBA Revision - Cloze Passage

Find a suitable word from the box for each of the blanks in the passage below.

mess

scolded

scraps

piece

floor

handful

allowed

dozens

out

bin

John's mother was furious with him yesterday. She was cleaning his room and found it in a huge _____. There were _____ of food under the bed and _____ of old sweet wrappers. There was a _____ full of rubbish which he had not thrown _____. There was even a _____ of mouldy old cake behind his desk. She almost slipped on a _____ of marbles on the _____. John was _____ severely and his mother told him if he did not keep his room _____, he would not be _____ to watch television any more.



Answers Key Cloze 4

Disha Experts

A red circular graphic with a gradient, appearing as a partial circle or a thick arrow pointing to the right, located to the right of the 'Disha Experts' text.

Answers Key Cloze 4:

The ELL Teacher's Toolbox Larry Ferlazzo, Katie Hull Sypnieski, 2018-04-24 Practical strategies to support your English language learners The ELL Teacher's Toolbox is a practical valuable resource to be used by teachers of English Language Learners in teacher education credential programs and by staff development professionals and coaches It provides hundreds of innovative and research based instructional strategies you can use to support all levels of English Language Learners Written by proven authors in the field the book is divided into two main sections Reading Writing and Speaking Listening Each of those sections includes Top Ten favorites and between 40 and 70 strategies that can be used as part of multiple lessons and across content areas Contains 60% new strategies Features ready to use lesson plans Includes reproducible handouts Offers technology integration ideas The percentage of public school students in the U S who are English language learners grows each year and with this book you ll get a ton of fresh innovative strategies to add to your teaching arsenal

The ELL Teacher's Toolbox 2.0 Larry Ferlazzo, Katie Hull Sypnieski, 2025-01-24 Set your English language learners up for success with this effective resource The ELL Teacher's Toolbox 2.0 is a valuable updated resource that teachers of English Language Learners ELLs can use to improve student outcomes With hundreds of innovative strategies and activities to bring to your classroom this book can be used with learners of all levels and in any instructional setting This revised edition provides the latest enhancements to the instructional tools along with 16 new chapters that you can add to your teaching repertoire New content includes coverage of artificial intelligence online learning environments and differentiated instruction Graphics and visuals make it easy to understand and adapt the content to your unique teaching situation Written by proven authors in the field the book is divided into two main sections Reading Writing and Speaking Listening Each of those sections includes Top Ten favorites and between 40 and 70 strategies that can be used as part of multiple lessons and across content areas Contains 60% new strategies Features ready to use lesson plans Includes reproducible handouts Offers technology integration ideas For the growing number of ELLs in public schools effective ELL instruction can mean the difference between long term academic success and continued struggling In this book you ll find countless practical ideas to add to your teaching arsenal or for training and coaching ELL teachers so you can support your students on their journey

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web limiting reactant theoretical and percentage yields key $4 \text{ko}_2 + 5 \text{h}_2 \rightarrow 2 \text{h}_2\text{o} + 4 \text{koh}$ aq 3 o_2 g if a reaction vessel contains 0 mol ko_2 and 0 mol h_2 o what is the limiting reactant

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the balanced equation shown below if the reaction of 90.6 grams of CO produces 36.7 grams of C₃H₈ what is the percent yield
 $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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web start with 8.3 mole H₂S and multiply it with 2 mole H₂O over 2 mole H₂S from the equation
 $8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$
 8.3 mole H₂O $\times$ 18g H₂O / 1 mole H₂O = 149.4g H₂O is your theoretical yield so percent yield = $\frac{137.1\text{g H}_2\text{O}}{149.4\text{g H}_2\text{O}} \times 100 = 91.77\%$
 problems 6 & 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH₃ can be produced theoretically from the reaction of 5.0 g of N
 what is the limiting reagent if 8.52 g are actually formed what is the

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 $21.3 \text{ g O}_2 \times \frac{1 \text{ mole O}_2}{32 \text{ g O}_2} = 0.6656 \text{ moles O}_2$

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web percent yield is a measure of the efficiency of the experimental design
 yield = $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$
 in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction Al + O₂ → Al₂O₃
 4 Al + 3 O₂ → 2 Al₂O₃
 b determine the theoretical yield of Al₂O₃
 $160.0 \text{ g Al} \times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g Al}_2\text{O}_3 = 151.96 \text{ g Al}_2\text{O}_3$
 302.3 g Al₂O₃ / 26.982 g Al = 4 mol Al = 1 mol Al₂O₃
 c determine the percent yield
 yield = $\frac{\text{actual}}{\text{theoretical}} \times 100 = \frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$

7.2 limiting reagent and reaction yields chem 1114 - Jul 02 2022

web 7 Na₂C₂O₄ is the limiting reactant percent yield = 86.68% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of H₂ are left over 13

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web the percent yield is calculated as follows text percent yield $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ percent yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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