

Pre-lab Questions for Experiment #5

Rates of Chemical Reactions I: A Clock Reaction

1) Complete and balance the following reactions:



2) The rate law for the reaction



is $\text{rate} = k[\text{H}_2][\text{NO}]^2$. Which of the following mechanisms can be ruled out on the basis of the observed rate expression?

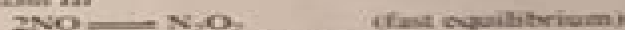
Mechanism I



Mechanism II



Mechanism III



3) Consider the reaction



From the following data obtained at a certain temperature, determine the order of the reaction and calculate the rate constant:

[A] (M)	[B] (M)	Rate (M/s)
1.50	1.50	3.20×10^{-1}
1.50	2.50	3.20×10^{-1}
3.00	1.50	6.40×10^{-1}

4) The following gas-phase reaction was studied at 290°C by observing the change in pressure as a function of time in a constant-volume vessel:



Determine the order of the reaction and the rate constant based on the following data:

Time (s)	Partial Pressure of $\text{ClCO}_2\text{CCl}_3$ (mmHg)
0	15.76
181	12.64
513	8.73
1164	4.44

Answer Of Pre Lab Rates Of Reaction

Raffaela Di Napoli



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Integrated Approach to Coordination Chemistry Rosemary A. Marusak, Kate Doan, Scott D. Cummings, 2007-03-07 Coordination chemistry is the study of compounds formed between metal ions and other neutral or negatively charged molecules This book offers a series of investigative inorganic laboratories approached through systematic coordination chemistry It not only highlights the key fundamental components of the coordination chemistry field it also exemplifies the historical development of concepts in the field In order to graduate as a chemistry major that fills the requirements of the American Chemical Society a student needs to take a laboratory course in inorganic chemistry Most professors who teach and inorganic chemistry laboratory prefer to emphasize coordination chemistry rather than attempting to cover all aspects of inorganic chemistry because it keeps the students focused on a cohesive part of inorganic chemistry which has applications in medicine the environment molecular biology organic synthesis and inorganic materials **Lab Manual Chemistry Class XII -by Dr. K. N. Sharma, Dr. Subhash Chandra Rastogi, Er. Meera Goyal (SBPD Publications)** Dr. K. N. Sharma, Dr. Subhash Chandra Rastogi, Er. Meera Goyal, 2021-07-03 Highly Useful for Various Engineering and Medical Competitive Examinations **Chemistry** Lemay, Prentice-Hall Staff, 1996 **Nuclear Science Abstracts**, 1971-05 **Chemical Investigations** Nancy Konigsberg Kerner, 1986 *The Latest and Best of TESS*, 1991 **Chemistry** McGraw-Hill Staff, 2001-07 **Experimental Organic Chemistry** Daniel R. Palleros, 2000-02-04 This cutting edge lab manual takes a multiscale approach presenting both micro semi micro and macroscale techniques The manual is easy to navigate with all relevant techniques found as they are needed Cutting edge subjects such as HPLC bioorganic chemistry multistep synthesis and more are presented in a clear and engaging fashion Who's the New Kid in Chemistry? John D. Butler, 2013-12-12 Who's the New Kid in Chemistry offers an unprecedented look at student engagement and teacher best practices through the eyes of an educational researcher enrolled as a public high school student Over the course of seventy nine consecutive days John D Butler participates in and observes Rhode Island 2013 Teacher of the Year Jessica M Waters's high school chemistry class documenting his experiences as they unfold Who's the New Kid in Chemistry is a compelling example of what can be accomplished when an educational researcher and teacher collaborate in the classroom This work includes a discussion on flexible homework assignments data driven instruction and thirty teacher best practices This book is an invaluable resource for teachers across all content areas masters and doctoral research method classes and future Teachers of the Year **Laboratory Manual on Biotechnology** P. M. Swamy, 2008 **Selected Water Resources Abstracts**, 1988-04 *Chemical Engineering Education*, 2002 Pre-lab Exercises for Experimental Organic Chemistry Royston M. Roberts, John C. Gilbert, Stephen F. Martin, 1994 **General Chemistry** Hunt, 1976-06-07 **Laboratory Experiments for General Chemistry** Harold R. Hunt, Toby F. Block, George M. McKelvy, 1998 Methods in Biotechnology Seung-Beom Hong, M. Bazlur Rashid, Lory Z. Santiago-Vázquez, 2016-08-01 As rapid advances in biotechnology occur there is a need for a

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Experimental Organic Chemistry John C. Gilbert, Stephen F. Martin, 2006 This proven and well tested laboratory manual for organic chemistry students contains procedures for both miniscale also known as small scale and microscale users This lab manual gives students all the necessary background to enter the laboratory with the knowledge to perform the experiments with confidence For the microscale labs experiments were chosen to provide tangible quantities of material which can then be analyzed Chapters 1 2 introduce students to the equipment record keeping and safety of the laboratory Chapters 3 6 and 8 are designed to introduce students to laboratory techniques needed to perform all experiments In Chapters 7 and 9 through 20 students are required to use the techniques to synthesize compounds and analyze their properties In Chapter 21 students are introduced to multi step syntheses of organic compounds a practice well known in chemical industry In Chapter 23 students are asked to solve structures of unknown compounds The new chapter 24 introduces a meaningful experiment into the textbook that reflects the increasing emphasis on bioorganic chemistry in the sophomore level organic lecture course This experiment not only gives students the opportunity to accomplish a mechanistically interesting and synthetically important coupling of two amino acids to produce a dipeptide but also provides valuable experience regarding the role of protecting groups in effecting synthetic transformations with multiple functionalized molecules

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