

PhET Isotopes and Atomic Mass

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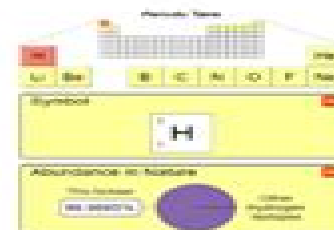
Date 10-2-2022

Use the link <https://phet.colorado.edu/en/simulation/isotopes-and-atomic-mass> to go to the PhET website for the simulation Isotopes and Atomic Mass.

Begin with the **Isotopes and Atomic Mass**.

Click on the + symbols next to the bar with Symbol and Abundance in Nature to display the data.

You can build isotopes of different atoms by clicking on the symbol and adding or subtracting neutrons using the bowl of neutrons next to the scale. You can also click on the scale to change to Atomic Mass (amu).



Protons: 5
Neutrons: 5
Electrons: 5

Let's practice with Boron.

1) What are the symbol, atomic mass and abundance in nature of the isotope Boron - 11?



Now add one neutron to make Boron -12. Notice that the atom shakes. This indicates that the isotope is unstable. Unstable isotopes undergo radioactive decay. You will need to note this for other elements in the table below.

Symbol	$^{13}_6\text{C}$	$^{14}_7\text{N}$	$^{18}_8\text{O}$	$^{10}_4\text{Be}$	^4_2He	$^{19}_9\text{F}$
Name	Carbon-13		oxygen-18	Beryllium-10	helium-4	
Protons	6	7	8	4	2	9
Neutrons	7	6	10	6	2	10
Electrons	6	7	8	4	2	9
Atomic Mass	13.00335	13	17.99916	10.01353	4.00260	18.99840
Abundance	1.07%	0	0.205%	trace	99.9999	100%
Stable OR Unstable	stable	unstable	stable	unstable	stable	stable

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Bishop, Roy Lovitt



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