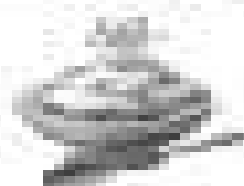


Atomic Probing Structure of the Atom

Name _____
Date _____



Purpose:

To determine the structure of the atom by observing the deflection of alpha particles.

Part I: Finding Evidence for Nuclear Structure

Alpha particles are positively charged and have a mass of about 4 atomic mass units. They are emitted from the nucleus of a radioactive substance. In the Rutherford experiment, alpha particles are directed at a thin foil of gold. The deflection of the particles is observed by a fluorescent screen.

Observation: Most alpha particles pass straight through the foil, but a small number are deflected at large angles.

This observation suggests that the atom is mostly empty space, with a small, dense, positively charged nucleus.

Conclusion: The results of the Rutherford experiment suggest that the atom has a small, dense, positively charged nucleus, and that most of the atom is empty space.

From this experiment, we can conclude that the atom is mostly empty space, with a small, dense, positively charged nucleus.

Observation: The deflection of alpha particles is proportional to the square of the atomic number of the foil.

This observation suggests that the nucleus is positively charged, and that the deflection is due to the repulsive force between the alpha particles and the nucleus.

Answers Lesson 11 Atomic Pudding

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