



Introduction to working as a Chemist.

Use a highlighter as you work through this booklet. ESL students, make vocab lists of the terminology that is defined or described here.

Assessment Statements

- | | |
|--|---|
| <ul style="list-style-type: none"> ☐ Outline the Scientific Method ☐ Identify independent, dependent & controlled variables. ☐ Demonstrate correct and safe use of laboratory equipment. ☐ Distinguish between quantitative and qualitative data <ul style="list-style-type: none"> ➤ <i>Qualitative is descriptive, not using numbers, e.g. properties, colour, gas production, sound, odor.</i> ➤ <i>Quantitative is mathematical, including</i> <ul style="list-style-type: none"> ◊ <i>Raw data (as recorded)</i> ◊ <i>Processed (after calculations)</i> ☐ Distinguish between accuracy and precision <ul style="list-style-type: none"> ➤ <i>Accuracy describes the correctness of a set of measurements</i> ➤ <i>Precision describes the repeatability of a set of measurements</i> | <ul style="list-style-type: none"> ☐ State SI units of measurement for <ul style="list-style-type: none"> ➤ <i>Mass: grams (g)</i> ➤ <i>Length: Metres (m), millimeters (mm)</i> ➤ <i>Concentration: molarity (M)</i> ➤ <i>Temperature: Degrees Celsius (°C)</i> ➤ <i>Volume: Litres (L) or millilitres (mL)</i> ➤ <i>Density: grams per cubic centimeter (g/cm³)</i> ➤ <i>Conductivity: Siemens (S)</i> ➤ <i>Acidity or alkalinity (pH)</i> ☐ Energy: Joules (J or kJ)Determine the uncertainty of digital and analog measuring tools <ul style="list-style-type: none"> ➤ <i>Usually ± 1 of the smallest division</i> ☐ Calculate the % error of a given or recorded set of values <ul style="list-style-type: none"> ➤ <i>$(\text{Measured} - \text{True}) / \text{True} \times 100$</i> ☐ State values to appropriate numbers of significant figures ☐ Calculate values to appropriate numbers of significant figures |
|--|---|

Related Concepts:

- ☐ 01.1 Element symbols (Quia quiz: <http://www.quia.com/quiz/3246640.htm#>)
- ☐ 01.2 Accuracy, Precision, Error, Uncertainty (Quia quiz: <http://www.quia.com/quiz/3507626.htm#>)
- ☐ 01.3 SigDigs and Notation (Quia quiz: <http://www.quia.com/quiz/3504378.htm#>)

Command Terms: Be sure you understand what is being asked of you

Analyse Break down to identify key parts. Interpret information to reach conclusions.

Apply	Use knowledge & understanding in response to a real situation or given circumstances.
Calculate	Obtain a numerical answer, showing relevant stages in the working.
Define	Give the precise meaning of a word, phrase, concept or physical quantity.
Describe	Give a detailed account or picture of an event, pattern, process or pattern.
Design	Produce a plan, simulation or model.
Determine	Obtain the only possible answer.
Discuss	Offer a considered and balanced review that includes a range of arguments, factors or hypotheses. Opinions or conclusions should be presented clearly and supported by appropriate evidence.
Distinguish	Make clear the differences between two or more concepts or items.
Evaluate	Assess the implications and limitations; make judgments about the ideas, works, solutions or ideas in relation to selected criteria.
Explain	Give a detailed account including reasons or causes.
Identify	Provide an answers from a number of possibilities. Recognise and state briefly a distinguishing fact or feature.
Label	Add title, labels or brief explanation(s) to a diagram or graph.
List	Give a sequence of answers with no explanation.
Outline	Give a brief account.
Recall	Remember or recognize from prior learning experiences.
State	Give a specific name, value or other brief answer without explanation or calculation.
Suggest/Propose	Propose a solution, hypothesis or other possible answer.

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Chemistry Karen Timberlake, 2003

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Introduction to Forensic Chemistry Kelly M. Elkins, 2018-09-03 Chemistry

Forensic Science Forensic chemistry is a subdiscipline of forensic science its principles guide the analyses performed in modern forensic laboratories Forensic chemistry s roots lie in medico legal investigation toxicology and microscopy and have since led the development of modern forensic analytic techniques and practices for use in a variety of applications Introduction to Forensic Chemistry is the perfect balance of testing methods and application Unlike other competing books on the market coverage is neither too simplistic nor overly advanced making the book ideal for use in both undergraduate and graduate courses The book introduces chemical tests spectroscopy advanced spectroscopy and chromatography to students The second half of the book addresses applications and methods to analyze and interpret controlled substances trace evidence questioned documents firearms explosives environmental contaminants toxins and other topics The book looks at innovations in the field over time including the latest development of new discernible chemical reactions instrumental tools methods and more Key features Nearly 300 full color figures illustrating key concepts and over 20 case studies Addresses all the essential topics without extraneous or overly advanced coverage Includes full pedagogy of chapter objectives key terms lab problems end of chapter questions and additional readings to emphasize key learning points Includes chemical structures and useful spectra as examples Fulfills the forensic chemistry course requirement in FEPAC

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