

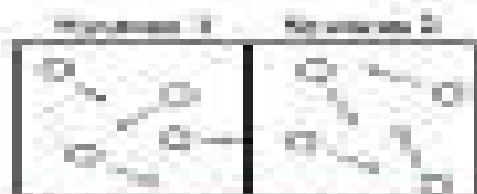
Unit 5: Thermochemistry**5.1: Heat Transfer and Thermal Equilibrium****5.2: Heat Capacity and Calorimetry****5.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving with v is $\frac{1}{2}mv^2$, so the more of particles that move fast enough, the more kinetic energy we will see. Temperature, therefore, is the thermally accessible kinetic energy that the particles possess as well as the distribution's maximum value to help describe the distribution's character.

3. Sketch each graph to represent the relative distribution of the particles in the following two gas samples at their same gas:



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.
 System 1 is at a lower temperature than System 2. Temperature is a measure of the average KE of the particles in the system. In System 1, the gas particles are closer together, and all the individual particles have the same mass, depending on their separation KE is the same, so all of the particles have the same mass, so they are moving faster than the particles in System 2. The particles in System 1 are closer together, so they are moving faster than the particles in System 2. The particles in System 1 are closer together, so they are moving faster than the particles in System 2.
- b. In which direction will kinetic energy be transferred? Justify your answer in terms of molecular collisions.
 KE will be transferred from System 1 to System 2. The particles in System 1 are closer together, so they are moving faster than the particles in System 2. The particles in System 1 are closer together, so they are moving faster than the particles in System 2. The particles in System 1 are closer together, so they are moving faster than the particles in System 2.

Answer Key Unit 6 Thermochemistry

Darrell D. Ebbing



Answer Key Unit 6 Thermochemistry:

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