

Lab 10: Calculations with a Chemical Structure

Objectives:

1. To derive an empirical formula from a chemical structure without performing stoichiometric calculations.
2. To calculate the number of moles of a solid from a weighing measurement in the lab.
3. To determine the molar mass of a solid.
4. To determine the theoretical amount of product that would be produced.
5. To compare the theoretical amount to the actual amount of substance formed and calculate the percent yield.

Reaction equation:

Reaction of sodium chloride solution: 25.0 mL

Reaction of calcium chloride solution: 25.0 mL

Reaction of silver nitrate: 5.00 g

Reaction of silver nitrate: 5.00 g

Reaction of silver nitrate: 5.00 g

Reaction of silver nitrate: 5.00 g

Procedure:

Calculations and Results:

1. Calculations:

$$\text{moles of sodium chloride} = 25.0 \text{ mL soln. } \times \frac{1.0 \text{ M}}{1000 \text{ mL}} = 0.0250 \text{ mol NaCl}$$

$$\text{moles of calcium chloride} = 25.0 \text{ mL soln. } \times \frac{1.0 \text{ M}}{1000 \text{ mL}} = 0.0250 \text{ mol CaCl}_2$$

$$\text{moles of silver nitrate produced} = 5.00 \text{ g} \times \frac{1 \text{ mol}}{169.87 \text{ g}} = 0.0295 \text{ mol AgNO}_3$$

$$\text{moles of product that would be produced} = 0.0250 \text{ mol CaCl}_2 \times \frac{2 \text{ mol}}{1 \text{ mol}} = 0.0500 \text{ mol CaCl}_2$$

2. Chemical equation: $\text{CaCl}_2 + 2\text{AgNO}_3 \rightarrow 2\text{AgCl} + \text{Ca(NO}_3)_2$

3. Reaction equation:

$$\text{Since } 0.0250 \text{ mol CaCl}_2 \times \frac{2 \text{ mol AgCl}}{1 \text{ mol CaCl}_2} = 0.0500 \text{ mol AgCl} \text{ (theoretical yield)}$$

$$\text{The } 0.0250 \text{ mol CaCl}_2 \text{ and theoretical yield is } 0.0500 \text{ mol}$$

$$\text{Percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100\% = \frac{0.0295 \text{ mol}}{0.0500 \text{ mol}} \times 100\% = 59.0\%$$

Answers For Lab 6b Heath Chemistry

L Reisser



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