

LAB MANUAL CHEMISTRY CLASS XII

Easy!

Experiment 1

Aim: To prepare 250 ml of 0.1N of Mohr's salt

Theory

Molecular mass of Mohr's salt

$\text{FeSO}_4 \cdot (\text{NH}_4)_2\text{SO}_4 \cdot 6\text{H}_2\text{O}$ is 392 g

Hence, for preparing 250 ml of 0.1N Mohr's salt

solution 3.92 g of salt is required

for preparing 250 ml of 0.1N Mohr's salt
solution $\frac{3.92 \times 250 \times 100}{1000} = 9.9 \text{ g}$ is required

Apparatus required: watch glass, weights, 250 ml beaker, glass rod, wash bottle.

Chemicals required: Mohr's salt, conc. H_2SO_4 ,
distilled water

Procedure

1. Record weight of clean dry watch glass

2. Weigh and record exactly 9.9 g of
Mohr's salt crystals

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 - The Rise of Digital Reading Answer To Hands On Lab Manual Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer To Hands On Lab Manual Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer To Hands On Lab Manual Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer To Hands On Lab Manual Chemistry
 - Personalized Recommendations
 - Answer To Hands On Lab Manual Chemistry User Reviews and Ratings
 - Answer To Hands On Lab Manual Chemistry and Bestseller Lists

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 - Answer To Hands On Lab Manual Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer To Hands On Lab Manual Chemistry
 - Highlighting and Note-Taking Answer To Hands On Lab Manual Chemistry
 - Interactive Elements Answer To Hands On Lab Manual Chemistry
8. Staying Engaged with Answer To Hands On Lab Manual Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer To Hands On Lab Manual Chemistry
9. Balancing eBooks and Physical Books Answer To Hands On Lab Manual Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer To Hands On Lab Manual Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answer To Hands On Lab Manual Chemistry
 - Setting Reading Goals Answer To Hands On Lab Manual Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer To Hands On Lab Manual Chemistry
 - Fact-Checking eBook Content of Answer To Hands On Lab Manual Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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