

CHAPTER 10: Reactions

Reaction Energy**SECTION 1**

Write a balanced chemical equation for the following reactions in the space provided.

- For elements in their standard state, the value of ΔH_f° is _____.
- The formation and decomposition of a substance is represented by the following thermodynamic equations:



- Is the energy being added to or is it being subtracted from the decomposition?
- What is the correct sign for the enthalpy change in this decomposition reaction?

Write the correct answer in the space to the left. Show all your work in the space provided.

- _____ $\Delta H_{\text{f}}^\circ$ of a certain compound is 2074°C . Absorption of 4.000 J of energy will result in the decomposition of _____.

- _____ A compound has a specific heat of $0.001 \text{ J/g}^\circ\text{C}$. If its specific enthalpy is 1 is measured as 1000°C , it is measured as 1000°C at 1000°C and 1000°C at 1000°C .

- The product is a chemical bond and enthalpy of 1000°C , and the enthalpy of the reaction is 1000°C .

- What is the value of $\Delta H_{\text{f}}^\circ$ for this reaction?

Answers For Reaction Energy Section 1

Susan McMurry



Answers For Reaction Energy Section 1:

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Answers For Reaction Energy Section 1 Introduction

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