# **HOMOGENEOUS AQUEOUS SYSTEMS**

# Section Review

#### Objectives

- · Distinguish between a solvent and a solute
- · Describe what happens in the solution process
- Explain why all ionic compounds are electrolytes
- Demonstrate how to write the formula for a hydrate

#### Vocabulary

- aqueous solution
- solvent.
- solute

- solvation.
- electrolyte
- monelectrolyte

- strong electrolyte
- weak electrolyte
- hydrate.

#### **Key Equation**

Percent H<sub>2</sub>O = mass of water x 100%

mass of hydrate

#### Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

Water is a polar liquid and an excellent \_\_\_\_\_\_ for many substances. Aqueous solutions are \_\_\_\_\_ mixtures of ions or molecules in water. The solubility of a solute depends on solute-

solvent interactions. A good rule to remember is \_\_\_\_3\_\_.

Substances that dissolve as ions are known as \_\_4\_\_\_ A
solute that is completely ionized in solution is a \_\_5\_\_
electrolyte. A weak electrolyte is only \_\_6\_\_ ionized. A solution
of an electrolyte will \_\_7\_\_ an electric current, whereas a

solution of a \_\_\_\_8 \_\_ is nonconducting.

Many crystals are \_\_\_\_\_\_; they contain water of hydration.

In the process called \_\_\_\_\_\_ the water of hydration is lost from a

hydrate that is exposed to air.

3.

4.

5. \_\_\_\_\_

6.

7.

8. \_\_\_\_

10

Chapter 15 Water and Aqueous Systems 379

# Answers For Homogeneous Aqueous Systems Section Review

**U.S. Nuclear Regulatory Commission. Office of Standards Development** 

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