

Stoichiometry

Solve the following problems:

1. Four moles of iron (II) sulfide added to seven moles of oxygen, produced two moles of iron (III) oxide and four moles of sulfur dioxide. (20pts)

a) Translate the problem to chemical equation.



Reactant side

4 Fe atoms

4 S atoms

7(2) =14 O atoms

Product side

4 Fe atoms

4 S atoms

2(3) +4(2) =14 O atoms

- b) How many grams of oxygen will react with 2.0 grams of iron (II) sulfide?

Given:

m FeS = 2.0 g

Req'd. = m O₂

Solution:

molar mass of FeS = 88 g/mol

mole ratio = 4 mol FeS: 7 mol O₂

molar mass of O₂ = 32 g/mol

$$m_{O_2} = 2.0 \cancel{g_{FeS}} \left(\frac{1 \cancel{mol_{FeS}}}{88 \cancel{g_{FeS}}} \right) \left(\frac{7 \cancel{mol_{O_2}}}{4 \cancel{mol_{FeS}}} \right) \left(\frac{32 \cancel{g_{O_2}}}{1 \cancel{mol_{O_2}}} \right) = 1.3 g_{O_2}$$

2. Rusting involves the reaction of iron, Fe with oxygen, O₂ to form iron (III) oxide, Fe₂O₃. (20pts)

a) Write the balanced chemical equation.



Reactant side

4 Fe atoms

3(2) =6 O atoms

Product side

4 Fe atoms

2(3) =6 O atoms

- b) What are the mole ratios of the product to each reactant and of the reactants to each other?

Reactant to each other:

4 mol Fe: 3 mol O₂

Product to each reactant:

2 mol Fe₂O₃:4 mol Fe

Answer Key To Stoichiometry Homework Problems

Leo J. Malone, Theodore O. Dolter



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Table of Contents Answer Key To Stoichiometry Homework Problems

1. Understanding the eBook Answer Key To Stoichiometry Homework Problems
 - The Rise of Digital Reading Answer Key To Stoichiometry Homework Problems
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key To Stoichiometry Homework Problems
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key To Stoichiometry Homework Problems
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key To Stoichiometry Homework Problems
 - Personalized Recommendations
 - Answer Key To Stoichiometry Homework Problems User Reviews and Ratings
 - Answer Key To Stoichiometry Homework Problems and Bestseller Lists
5. Accessing Answer Key To Stoichiometry Homework Problems Free and Paid eBooks
 - Answer Key To Stoichiometry Homework Problems Public Domain eBooks
 - Answer Key To Stoichiometry Homework Problems eBook Subscription Services
 - Answer Key To Stoichiometry Homework Problems Budget-Friendly Options

6. Navigating Answer Key To Stoichiometry Homework Problems eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Key To Stoichiometry Homework Problems Compatibility with Devices
 - Answer Key To Stoichiometry Homework Problems Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key To Stoichiometry Homework Problems
 - Highlighting and Note-Taking Answer Key To Stoichiometry Homework Problems
 - Interactive Elements Answer Key To Stoichiometry Homework Problems
8. Staying Engaged with Answer Key To Stoichiometry Homework Problems
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key To Stoichiometry Homework Problems
9. Balancing eBooks and Physical Books Answer Key To Stoichiometry Homework Problems
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key To Stoichiometry Homework Problems
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answer Key To Stoichiometry Homework Problems
 - Setting Reading Goals Answer Key To Stoichiometry Homework Problems
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Key To Stoichiometry Homework Problems
 - Fact-Checking eBook Content of Answer Key To Stoichiometry Homework Problems
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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