

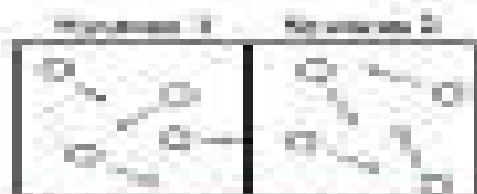
Unit 6: Thermochemistry**6.1: Heat Transfer and Thermal Equilibrium****6.2: Heat Capacity and Calorimetry****6.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving with v is $\frac{1}{2}mv^2$, so the more of particles that move fast enough, the more kinetic energy we will see for a system. Therefore, the temperature of a system is proportional to the average kinetic energy of the particles. The Boltzmann-probability distribution curves show that the average kinetic energy of all of the particles is proportional to the temperature of the system.

3. Sketch a curve meant to represent the relative distribution of the particles in the following two gases, samples of the same gas.



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.
System 1 is at a lower temperature than System 2. Temperature is a measure of the average KE of the particles in the system. In System 1, the particles are more spread out than in System 2, so the particles in System 1 have more kinetic energy. Therefore, the temperature of System 1 is higher than the temperature of System 2.
- b. In which direction will thermal energy be transferred? Justify your answer in terms of molecular collisions.
Thermal energy will be transferred from System 1 to System 2. The particles in System 1 have more kinetic energy than the particles in System 2. When the particles from the two systems are mixed, the particles from System 1 will collide with the particles from System 2, and the particles from System 1 will transfer some of their kinetic energy to the particles from System 2. This will continue until the two systems reach thermal equilibrium.

Answer Key Unit 6 Thermochemistry

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