

Mole Calculations Review Worksheet – answers on next page.

Complete the problems on this worksheet to review for the Chapter 10 test. Be sure to use dimensional analysis, and label all numbers with units when solving the problems.

- Calculate the molar mass of each compound.
 - LiOH
 - barium bromide
 - $\text{Mg}(\text{C}_2\text{H}_3\text{O}_2)_2$
 - $\text{Ca}(\text{NO}_3)_2$
- How many molecules are in 45.0 grams CH_4 ?
- How many moles are in 18.8 grams NaOH?
- A salt shaker containing 9.58×10^{23} formula units NaCl contains how many moles?
- Which of the following has the greatest mass?
4.2 moles C or 8.34×10^{28} f.u. CaCO_3 or 12.6 g $\text{Al}(\text{NO}_3)_3$
- Find the number of representative particles in each of the following:
 - 0.28 mol Ni
 - 1.84 mol S
 - 1.32 mol F_2
 - 3.15 mol K_3PO_4
- How many liters would 1.45 moles oxygen gas occupy at STP?
- How many grams would 15.8 L of ammonia occupy at STP?
- How many total atoms are in 5.8 mol CaCl_2 ?
- Calculate the percent composition of each element in the following compounds:
 - iron(III) oxide
 - $\text{Ca}_3(\text{PO}_4)_2$
 - methane
 - H_2SO_4
- Calculate the empirical and molecular formulas of adrenaline. The compound has a molecular mass of 183 g/mol, and is 59.0% C, 7.1% H, 26.2% O, and 7.7% N by mass.
- A compound is 72.2% Mg and 27.8% N by mass. Calculate the empirical formula.
- Nicotine is 74.1% C, 8.6% H, and 17.3% N by mass. The molecular mass of nicotine is 162 g/mol. Calculate its empirical and molecular formulas.
- How many total ions are in 2.95 grams MgSO_4 ?

Answers Chemistry The Mole Review Questions

Julia Burdge



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Chemistry Neil D. Jespersen, Alison Hyslop, 2021-11-02 Chemistry The Molecular Nature of Matter 8th Edition continues to focus on the intimate relationship that exists between structure at the atomic molecular level and the observable macroscopic properties of matter. Key revisions in this edition focus on three areas: The deliberate inclusion of more updated real world examples that relate common real world student experiences to the science of chemistry; Simultaneously, examples and questions have been updated to align them with career concepts relevant to the environmental, engineering, biological, pharmaceutical, and medical sciences; Providing students with transferable skills with a focus on integrating metacognition and three dimensional learning into the text. When students know what they know, they are better able to learn and incorporate the material. Providing a total solution through New WileyPLUS by fully integrating the enhanced eText with online assessment, answer specific responses, and additional practice resources. The 8th edition continues to emphasize the importance of applying concepts to problem solving to achieve high level learning and increase retention of chemistry knowledge. Problems are arranged in an intuitive confidence building order.

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