
Chem - Covalent Bonding: Questions & Answers

properties of non metals ✓ Ans - inability to conduct electricity - absence of ions or delocalised electrons
low boiling and melting points - generally weak intermolecular bonds, can be soft

molecules ✓ Ans - separate groups of atoms with a formula, bonded together

intermolecular bonds ✓ Ans - bonds between molecules (dispersion, dipole-dipole, hydrogen bonding) determines physical properties

intramolecular bonds ✓ Ans - bonds between atoms in a molecule (covalent, ionic eg.)

covalent bonds ✓ Ans - non metals share electrons to complete the octet rule, an atom forms as many bonds as it needs eg. carbon needs 4 electrons so forms 4 bonds

why do non metals share electrons not transfer? ✓ Ans - non metals have a relatively high electronegativity so competing to attract each others electron, shares instead

electron dot diagrams ✓ Ans - lewis diagram shows valence electrons (bonding pairs) and lone pairs

valence structure ✓ Ans - each pair of bonding electrons is shown as a line, still represents dots as lone pairs. 1 line=2 electrons

bonding pairs and lone pairs ✓ Ans - bonding pairs; shared electrons in bond
lone pairs; non bonding pairs

double covalent bond ✓ Ans - rigid, two pairs of electrons (4 electrons) are shared, represented as a double line in valence structures eg. O_2

Answers For Chem Think Covalent Bonding

Victor M. Corman



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