

Lab 10: Calculations with a Chemical Structure

Objectives:

1. To derive an empirical formula from a chemical structure without performing stoichiometric calculations.
2. To calculate the number of moles of a solid from a weighing measurement in the lab.
3. To determine the molar mass of a solid.
4. To determine the theoretical amount of product that would be produced.
5. To compare the theoretical amount to the actual amount of substance formed and calculate the percent yield.

Reaction equation:

Reaction of sodium chloride solution: 25.0 mL

Reaction of calcium chloride solution: 25.0 mL

Reaction 10% Silver Paper: 0.00 g

Reaction 10% Silver Paper + Dry Solid: 0.00 g

Reaction of silver chloride solution: 25.0 mL

Reaction of silver chloride solution: 25.0 mL

Reaction of silver chloride solution: 25.0 mL

Procedure:

Reaction and Calculations:

1. Reaction:

$$\text{mole of sodium chloride} = 0.00 \text{ mol NaCl} \times \frac{1.0}{100.0} = 0.0000 \text{ mol NaCl}$$

$$\text{mole of calcium chloride} = 0.00 \text{ mol CaCl}_2 \times \frac{1.0}{100.0} = 0.0000 \text{ mol CaCl}_2$$

$$\text{mole of silver chloride produced} = 0.00 \text{ mol AgCl} \times \frac{1.0}{100.0} = 0.0000 \text{ mol AgCl}$$

$$\text{mole of silver chloride produced} = 0.00 \text{ mol AgCl} \times \frac{1.0}{100.0} = 0.0000 \text{ mol AgCl}$$

2. Reaction equation: $\text{Na}_2\text{CO}_3 + \text{CaCl}_2 \rightarrow \text{CaCO}_3 + 2\text{NaCl}$

3. Reaction equation:

$$\text{mole of } \text{CaCl}_2 = 0.00 \text{ mol} \times \frac{1.0}{100.0} = 0.0000 \text{ mol CaCl}_2$$

$$\text{mole of } \text{CaCl}_2 = 0.00 \text{ mol} \times \frac{1.0}{100.0} = 0.0000 \text{ mol CaCl}_2$$

$$\text{mole of } \text{CaCl}_2 = 0.00 \text{ mol} \times \frac{1.0}{100.0} = 0.0000 \text{ mol CaCl}_2$$

Answers For Lab 6b Heath Chemistry

R Bogdan



Answers For Lab 6b Heath Chemistry:

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