

## Researched expected results

### Test 1

|                                             | Magnesium chloride                                                                  | Calcium bromide                                                                     | Strontium chloride                                                                  | Barium chloride                                                         |
|---------------------------------------------|-------------------------------------------------------------------------------------|-------------------------------------------------------------------------------------|-------------------------------------------------------------------------------------|-------------------------------------------------------------------------|
| Initial                                     | Colourless solution                                                                 | Colourless solution                                                                 | Colourless solution                                                                 | Colourless solution                                                     |
| After adding 10 drops of 0.5M NaOH solution | Slightly white precipitate                                                          | Slightly white precipitate                                                          | Slightly white precipitate                                                          | Colourless solution                                                     |
| After adding excess NaOH                    | White precipitate                                                                   | Slightly white precipitate                                                          | Slightly white precipitate                                                          | Colourless solution                                                     |
| Reaction and Product formed                 | $\text{MgCl}_2 + 2\text{NaOH} \rightarrow \text{Mg(OH)}_2(\text{s}) + 2\text{NaCl}$ | $\text{CaBr}_2 + 2\text{NaOH} \rightarrow \text{Ca(OH)}_2(\text{s}) + 2\text{NaBr}$ | $\text{SrCl}_2 + 2\text{NaOH} \rightarrow \text{Sr(OH)}_2(\text{s}) + 2\text{NaCl}$ | $\text{BaCl}_2 + \text{NaOH} \rightarrow \text{Ba(OH)}_2 + \text{NaCl}$ |

### Test 2

|                                                                       | Magnesium chloride                                                              | Calcium bromide                                                                                                              | Strontium chloride                                                                              | Barium chloride                                                                          |
|-----------------------------------------------------------------------|---------------------------------------------------------------------------------|------------------------------------------------------------------------------------------------------------------------------|-------------------------------------------------------------------------------------------------|------------------------------------------------------------------------------------------|
| Initial                                                               | Colourless solution                                                             | Colourless solution                                                                                                          | Colourless solution                                                                             | Colourless solution                                                                      |
| After adding 10 drops of 1.0M H <sub>2</sub> SO <sub>4</sub> solution | Slight white precipitate                                                        | Slight white precipitate                                                                                                     | white precipitate                                                                               | white precipitate                                                                        |
| After adding excess H <sub>2</sub> SO <sub>4</sub>                    | Colourless solution                                                             | Slight white precipitate                                                                                                     | white precipitate                                                                               | white precipitate                                                                        |
| Reaction and Product formed                                           | $\text{MgCl}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{MgSO}_4 + 2\text{HCl}$ | $\text{CaBr}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{H}_2\text{O} + \text{SO}_2 + \text{Br}_2 + \text{CaSO}_4(\text{s})$ | $\text{SrCl}_2 + 2\text{H}_2\text{SO}_4 \rightarrow \text{Sr(HSO}_4)_2(\text{s}) + 2\text{HCl}$ | $\text{BaCl}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{BaSO}_4(\text{s}) + \text{HCl}$ |

### Test 3

|                                                                                                                                                                                            |                                                                                                                                                                                      |
|--------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|--------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|
| Observation of damp red litmus paper being held over the test tube containing a heated mixture of 10 drops of 0.1M ammonium chloride solution and 10 drops 0.5M sodium hydroxide solution. | Litmus paper turns blue                                                                                                                                                              |
| Reaction and product formed                                                                                                                                                                | $\text{NH}_4\text{Cl} + \text{NaOH} \rightarrow \text{NH}_3 + \text{NaCl} + \text{H}_2\text{O}$                                                                                      |
| Reason                                                                                                                                                                                     | Ammonium ions can be recognised in a solution. If ammonium ions are present when added to sodium hydroxide, they will be converted to ammonia gas and/or turn red litmus paper blue. |

### Test 4

|                                                                                                                                                                                      |                                  |
|--------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|----------------------------------|
| Observation of red litmus paper with the fumes of 1cm <sup>3</sup> solution of 0.5M sodium hydroxide                                                                                 | Turns damp red litmus paper blue |
| Observations made after a few minutes of dampening red litmus paper with deionised water and placing it beside filter paper soaked in 1M ammonia solution inside a closed petri dish | Turns damp red litmus paper blue |

### Test 5

| Solution                                                                                         | Volume (cm <sup>3</sup> ) | Molarity (M) | Observation                                                                                                        |
|--------------------------------------------------------------------------------------------------|---------------------------|--------------|--------------------------------------------------------------------------------------------------------------------|
| Limewater                                                                                        | 1                         | 0.02         | Colourless solution                                                                                                |
| $\text{HCl} + \text{Na}_2\text{CO}_3 \rightarrow \text{NaCl} + \text{H}_2\text{O} + \text{CO}_2$ | 3                         | 1.00         | After adding hydrochloric acid and sodium carbonate to the limewater, the solution went from colourless to cloudy. |

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**SJ Ball**



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## **Table of Contents A2 Chemistry 2013 Unofficial Mark Scheme**

1. Understanding the eBook A2 Chemistry 2013 Unofficial Mark Scheme
  - The Rise of Digital Reading A2 Chemistry 2013 Unofficial Mark Scheme
  - Advantages of eBooks Over Traditional Books
2. Identifying A2 Chemistry 2013 Unofficial Mark Scheme
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an A2 Chemistry 2013 Unofficial Mark Scheme
  - User-Friendly Interface
4. Exploring eBook Recommendations from A2 Chemistry 2013 Unofficial Mark Scheme
  - Personalized Recommendations
  - A2 Chemistry 2013 Unofficial Mark Scheme User Reviews and Ratings
  - A2 Chemistry 2013 Unofficial Mark Scheme and Bestseller Lists
5. Accessing A2 Chemistry 2013 Unofficial Mark Scheme Free and Paid eBooks

- A2 Chemistry 2013 Unofficial Mark Scheme Public Domain eBooks
- A2 Chemistry 2013 Unofficial Mark Scheme eBook Subscription Services
- A2 Chemistry 2013 Unofficial Mark Scheme Budget-Friendly Options
- 6. Navigating A2 Chemistry 2013 Unofficial Mark Scheme eBook Formats
  - ePub, PDF, MOBI, and More
  - A2 Chemistry 2013 Unofficial Mark Scheme Compatibility with Devices
  - A2 Chemistry 2013 Unofficial Mark Scheme Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of A2 Chemistry 2013 Unofficial Mark Scheme
  - Highlighting and Note-Taking A2 Chemistry 2013 Unofficial Mark Scheme
  - Interactive Elements A2 Chemistry 2013 Unofficial Mark Scheme
- 8. Staying Engaged with A2 Chemistry 2013 Unofficial Mark Scheme
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers A2 Chemistry 2013 Unofficial Mark Scheme
- 9. Balancing eBooks and Physical Books A2 Chemistry 2013 Unofficial Mark Scheme
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection A2 Chemistry 2013 Unofficial Mark Scheme
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine A2 Chemistry 2013 Unofficial Mark Scheme
  - Setting Reading Goals A2 Chemistry 2013 Unofficial Mark Scheme
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of A2 Chemistry 2013 Unofficial Mark Scheme
  - Fact-Checking eBook Content of A2 Chemistry 2013 Unofficial Mark Scheme
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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