
AAMC MCAT Practice Exam 1 Questions and Answers

In the chromatography of the reaction mixture, water absorbed on cellulose functioned as the stationary phase. What was the principal factor determining the migration of individual components in the sample?

- A) Hydrogen bonding
 - B) Solute concentration
 - C) Stationary phase concentration
 - D) Thickness of paper
- Answer - A) Hydrogen bonding

The answer to this question is A because the relative amount of hydrogen bonding to the stationary phase will determine the relative rate of migration of the various components in the sample.

What assumption is being made if scientists conclude that aspartic acid was formed by the prebiological synthesis in the passage?

- A) Aspartic acid is unstable at temperatures below 150 degrees C.
 - B) All of the malic acid underwent the dehydration reaction to form fumaric/maleic acid.
 - C) Compound A and cyanide were available on primitive Earth.
 - D) The reaction between ammonia and fumaric acid was catalyzed by the presence of water.
- Answer - C) Compound A and cyanide were available on primitive Earth.

The answer to this question is C since, in order for the experimental reaction sequence to be relevant to the primordial formation of aspartic acid, the starting materials used (Compound A and cyanide) are assumed to have been available.

According to the developed chromatography plate shown below, what is the approximate R_f value of aspartic acid?

aspartic acid = 2
solvent front = 10

- A) 0.20
 - B) 0.50
 - C) 5
 - D) 10
- Answer - The answer to this question is A (0.20) because R_f is the ratio of the distance travelled by the analyte relative to the solvent front during a chromatographic separation. Aspartic acid travelled two units, while the solvent front travelled ten units, giving an R_f of $2/10 = 0.20$ for aspartic acid.

Which of the following statements does NOT correctly describe the

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K Morrison

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Table of Contents Aamc Mcat Practice Test 11

1. Understanding the eBook Aamc Mcat Practice Test 11
 - The Rise of Digital Reading Aamc Mcat Practice Test 11
 - Advantages of eBooks Over Traditional Books
2. Identifying Aamc Mcat Practice Test 11
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Aamc Mcat Practice Test 11
 - User-Friendly Interface
4. Exploring eBook Recommendations from Aamc Mcat Practice Test 11
 - Personalized Recommendations
 - Aamc Mcat Practice Test 11 User Reviews and Ratings

- Aamc Mcat Practice Test 11 and Bestseller Lists
- 5. Accessing Aamc Mcat Practice Test 11 Free and Paid eBooks
 - Aamc Mcat Practice Test 11 Public Domain eBooks
 - Aamc Mcat Practice Test 11 eBook Subscription Services
 - Aamc Mcat Practice Test 11 Budget-Friendly Options
- 6. Navigating Aamc Mcat Practice Test 11 eBook Formats
 - ePub, PDF, MOBI, and More
 - Aamc Mcat Practice Test 11 Compatibility with Devices
 - Aamc Mcat Practice Test 11 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Aamc Mcat Practice Test 11
 - Highlighting and Note-Taking Aamc Mcat Practice Test 11
 - Interactive Elements Aamc Mcat Practice Test 11
- 8. Staying Engaged with Aamc Mcat Practice Test 11
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Aamc Mcat Practice Test 11
- 9. Balancing eBooks and Physical Books Aamc Mcat Practice Test 11
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Aamc Mcat Practice Test 11
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Aamc Mcat Practice Test 11
 - Setting Reading Goals Aamc Mcat Practice Test 11
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Aamc Mcat Practice Test 11
 - Fact-Checking eBook Content of Aamc Mcat Practice Test 11
 - Distinguishing Credible Sources

-
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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