



1987 Pontiac Sunbird Owners Manual

RC Schank



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web oxidation reduction redox reactions oxidation $\text{Cu} \rightarrow \text{Cu}^{2+} + 2\text{e}^{-}$ reduction $\text{MnO}_4^{-} + 4\text{H}^{+} + 3\text{e}^{-} \rightarrow \text{MnO}_2 + 2\text{H}_2\text{O}$ l when the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients $\text{Cu} : \text{MnO}_4^{-}$ is

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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^{-} + \text{Au} + \text{CN}^{-} + 2\text{e}^{-} + 2\text{H}^{+} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{HCN}$ make electrons equal convert to basic solution $4\text{CN}^{-} + 2\text{Au} + 2\text{CN}^{-} + 2\text{e}^{-}$ multiplied by a factor of 2 $2\text{e}^{-} + 2\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{H}^{+}$ 3 the final answer

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance ce o by adding ce h₂o

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation ni s ni a 2 a q 2 e now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balance each half reaction in basic solution 4 5 6 7 $\text{Cr}_2\text{O}_7^{2-}$ 2 NO_3^- 2 MnO_2 Cr^{3+} NO_3^- SO_2 Mn_2O_3 balance each redox reaction in acid solution using the half reaction method 8 H_2O_2 $\text{Cr}_2\text{O}_7^{2-}$ 2 9 TeO_3 2 N_2O_4 O_2 Cr^{3+} 10 ReO_4^- 11 PbO_2 IO_3^- I_2 Pb^{2+} 12 AsO_3^{3-} IO_3^- ReO_4^- H_2AsO_4^-

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web aug 29 2023 some points to remember when balancing redox reactions the equation is separated into two half equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding H^+ H_2O and e^- in this order balance the atoms in the equation apart from o and h

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively $\text{C}_6\text{H}_8\text{O}_6$ $2\text{H}_2\text{NO}_2$ $\text{C}_6\text{H}_6\text{O}_6$ $2\text{H}_2\text{O}_2$ choose 1 answer $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 a $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 b $\text{C}_6\text{H}_8\text{O}_6$ and H_2O_2 c $\text{C}_6\text{H}_8\text{O}_6$ and H_2O_2 d $\text{C}_6\text{H}_8\text{O}_6$ and H_2O_2

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$$\text{Pb(NO}_3)_2(\text{aq}) + 2\text{HCl}(\text{aq}) \rightarrow \text{PbCl}_2(\text{s}) + 2\text{HNO}_3$$

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web the net ionic equation for equation 1 is given as equation 3 s2 aq $\text{Fe}^{2+} \text{ aq}$ $\text{OH}^- \text{ aq}$ $\text{Fe}(\text{OH})_2 \text{ s}$ eq 3 8 note that the phase indicator aq is optional on ionic species but is

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web $\text{Cu}(\text{ClO}_4)_2$ is soluble all perchlorates are soluble while PbSO_4 is insoluble sulfates are generally soluble except for Pb^{2+} Ba^{2+} Hg_2^{2+} and Sr^{2+} the net ionic equation is

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