

Question 1

1(a)(iii) Difference: With powdered CaCO_3 the graph has a steeper slope at the start of the reaction than with marble chips because the rate of reaction is faster due to increased surface area of the solid.
Similarity: The total volume of CO_2 produced is the same in both cases since the same mass (and therefore number of moles) of CaCO_3 was used in both experiments.



1(a)(v) Moles of CaCO_3 used = $(1.8/100)$ mol. = 0.018 mol.
From the balanced equation for the reaction, 1 mole CaCO_3 produces 1 mol CO_2 .
Hence volume of gas produced at R.T.P. = (0.018×24000) cm^3 = 432 cm^3 .

1(a)(vi) There were impurities in the CaCO_3 used.

1(a)(vii) Concentration of the acid used and temperature at which the reaction was carried out.

1(b)(i) Fe^{3+} present.

**1(b)(ii) Magnesium ribbon decreases in size/dissolves.
Solution becomes pale green.**



1(b)(iv) SO_4^{2-} , SO_3^{2-} and CO_3^{2-} absent.

1(b)(v) White precipitate, which is soluble in aqueous ammonia, is formed.

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Sulthan Bathery Kerala in 2002 He Stood First in his College for B sc and M sc Chemistry He received his Ph D in Inorganic Chemistry from IIT Madras in the year 2010 His research interest includes phosphorus based ligands in synthetic inorganic chemistry and organometallic chemistry He has Published four research papers in reputed national and international journals He has more than four years of teaching experience in various engineering colleges **The Air Force Law Review** ,2015

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Table of Contents Answers For January 2014 Chemistry

1. Understanding the eBook Answers For January 2014 Chemistry
 - The Rise of Digital Reading Answers For January 2014 Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers For January 2014 Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers For January 2014 Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers For January 2014 Chemistry
 - Personalized Recommendations
 - Answers For January 2014 Chemistry User Reviews and Ratings
 - Answers For January 2014 Chemistry and Bestseller Lists
5. Accessing Answers For January 2014 Chemistry Free and Paid eBooks

- Answers For January 2014 Chemistry Public Domain eBooks
- Answers For January 2014 Chemistry eBook Subscription Services
- Answers For January 2014 Chemistry Budget-Friendly Options
- 6. Navigating Answers For January 2014 Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers For January 2014 Chemistry Compatibility with Devices
 - Answers For January 2014 Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers For January 2014 Chemistry
 - Highlighting and Note-Taking Answers For January 2014 Chemistry
 - Interactive Elements Answers For January 2014 Chemistry
- 8. Staying Engaged with Answers For January 2014 Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers For January 2014 Chemistry
- 9. Balancing eBooks and Physical Books Answers For January 2014 Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers For January 2014 Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers For January 2014 Chemistry
 - Setting Reading Goals Answers For January 2014 Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers For January 2014 Chemistry
 - Fact-Checking eBook Content of Answers For January 2014 Chemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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