

PHYSICAL KINETIC MOLECULAR THEORY

Learning Objectives

- 1) Identify theoretical differences in how gases participate in the kinetic theory and how molecules in liquids and solids participate in the kinetic molecular theory
- 2) Identify how temperature affects molecular motion
- 3) Explain how kinetic molecular theory is applied to the behavior of everyday substances

Module 1: Representation of States in Different Phases

The following figure illustrates representative particle behavior across different temperatures:



Figure 1: Representative particle behavior at different temperatures

Understanding the differences in particle behavior across different states is crucial for explaining various physical properties and processes. This module explores these differences in detail.

Key Concepts

1. **States of Matter:** The three primary states of matter are solid, liquid, and gas. Each state has distinct characteristics based on the arrangement and movement of its particles (see Figure 1). Understanding these states is essential for explaining the behavior of substances under different conditions.

Table 1: Characteristics of the Three States of Matter

	SOLID STATE	LIQUID STATE	GASEOUS STATE
Particle Arrangement and Movement:	Particles are packed closely together in a fixed, ordered arrangement. They vibrate in place but do not move from their positions.	Particles are in a disordered arrangement. They can move past one another, but they are still close together.	Particles are in a highly disordered, random arrangement. They move rapidly in all directions, filling the container.
Volume and Shape:	Solids have a fixed volume and a fixed shape. They maintain their shape regardless of the container.	Liquids have a fixed volume but take the shape of the container. They flow and adapt to the shape of the vessel.	Gases have no fixed volume and no fixed shape. They expand to fill the entire container and take its shape.

Answers Kinetic Molecular Theory Pogil

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