

Representing and interpreting nuclear reactions

We can represent nuclear reactions in the form of equations such as ${}^{14}_6\text{C} \rightarrow {}^{14}_7\text{N} + {}^0_{-1}\beta$

This equation represents the decay of a carbon-14 nucleus leaving a daughter nitrogen nucleus and emitting a beta particle in the process. Notice that the equation is 'balanced' in the sense that the total mass number and atomic number is the same on both sides of the arrow. This must always be the case. We know that energy is conserved in the reaction, so the energy released when a carbon-14 nucleus decays to a nitrogen nucleus is emitted in the form of a beta particle.

Worked example 8.9

Complete the following equation which represents a nuclear fission reaction and find the value of X.



The equation is balanced on both sides so the total mass number on the left side must be the same as the total mass number on the right side

$$235 + 1 = X + 139 + 2$$

$$236 = X + 141$$

$$236 - 141 = X$$

$$= 95$$

We can combine nuclear reaction equations and data about the mass of the elements on both sides of the equation to find out how much energy is released per fission.

Worked example 8.10

Calculate the energy released in the following fission reaction. Give your answer in MeV.



The data you need are:

mass of U-235 is 235.0439 u

mass of Ba-141 is 140.9144 u

mass of Kr-92 is 91.9262 u

mass of ${}^1_0\text{n}$ is 1.008 665 u.

Consider mass on each side of the equation

$$235.0439 \text{ u} + 1.008 66 \text{ u} \rightarrow 140.9144 \text{ u} + 91.9262 \text{ u} + (3 \times 1.008 665 \text{ u})$$

$$236.05256 \text{ u} \rightarrow 235.866595 \text{ u}$$

$$\Delta m = (236.05256 - 235.866595) \text{ u} \\ = 0.185 965 \text{ u}$$

$$\text{Energy released} = 0.185 96 \times 931.5 \text{ MeV} \\ = 173 \text{ MeV}$$

Albra 1 Page 339 Answers

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Table of Contents Alrbra 1 Page 339 Answers

1. Understanding the eBook Alrbra 1 Page 339 Answers
 - The Rise of Digital Reading Alrbra 1 Page 339 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Alrbra 1 Page 339 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Alrbra 1 Page 339 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Alrbra 1 Page 339 Answers
 - Personalized Recommendations
 - Alrbra 1 Page 339 Answers User Reviews and Ratings
 - Alrbra 1 Page 339 Answers and Bestseller Lists
5. Accessing Alrbra 1 Page 339 Answers Free and Paid eBooks
 - Alrbra 1 Page 339 Answers Public Domain eBooks
 - Alrbra 1 Page 339 Answers eBook Subscription Services
 - Alrbra 1 Page 339 Answers Budget-Friendly Options
6. Navigating Alrbra 1 Page 339 Answers eBook Formats

- ePub, PDF, MOBI, and More
- Alrbra 1 Page 339 Answers Compatibility with Devices
- Alrbra 1 Page 339 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Alrbra 1 Page 339 Answers
 - Highlighting and Note-Taking Alrbra 1 Page 339 Answers
 - Interactive Elements Alrbra 1 Page 339 Answers
- 8. Staying Engaged with Alrbra 1 Page 339 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Alrbra 1 Page 339 Answers
- 9. Balancing eBooks and Physical Books Alrbra 1 Page 339 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Alrbra 1 Page 339 Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Alrbra 1 Page 339 Answers
 - Setting Reading Goals Alrbra 1 Page 339 Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Alrbra 1 Page 339 Answers
 - Fact-Checking eBook Content of Alrbra 1 Page 339 Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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