

1989 Advanced Placement Exam

Part I Multiple Choice

Note: For all questions referring to solutions, assume that the solvent is water unless otherwise stated.

Directions: Each set of lettered choices below refers to the numbered statements immediately following it. Select the one lettered choice that best fits each statement and then blacken the corresponding space on the answer sheet. A choice may be used once, more than once, or not at all in each set.

Questions 1–3

- a. O c. Rb e. N
b. La d. Mg

- What is the most electronegative element?
- Which element exhibits the greatest number of different oxidation states?
- Which of the elements above has the smallest ionic radius for its most commonly found ion?

Questions 4–7

- a. $1s^2 2s^2 2p^5 3s^2 3p^5$
b. $1s^2 2s^2 2p^6 3s^2 3p^6$
c. $1s^2 2s^2 2p^6 2d^{10} 3s^2 3p^6$
d. $1s^2 2s^2 2p^6 3s^2 3p^6 3d^5$
e. $1s^2 2s^2 2p^6 3s^2 3p^6 3d^3 4s^2$

- An impossible electronic configuration
- The ground-state configuration for the atoms of a transition element
- The ground-state configuration of a negative ion of a halogen
- The ground-state configuration of a common ion of an alkaline earth element

Questions 8–10

- a. a solution with a pH less than 7 that is not a buffer solution
b. a buffer solution with a pH between 4 and 7
c. a buffer solution with a pH between 7 and 10
d. a solution with a pH greater than 7 that is not a buffer solution
e. a solution with a pH of 7

(Ionization constants:



- A solution prepared to be initially 1 M in NaCl and 1 M in HCl.
- A solution prepared to be initially 1 M in Na_2CO_3 and 1 M in CH_3COONa
- A solution prepared to be initially 0.5 M in CH_3COOH and 1 M in CH_3COONa

Questions 11–14

- a. hydrogen bonding
b. hybridization
c. ionic bonding
d. resonance
e. van der Waals forces (London dispersion forces)

- Is used to explain why iodine molecules are held together in the solid state
- Is used to explain why the boiling point of HF is greater than the boiling point of HBr
- Is used to explain the fact that the four bonds in methane are equivalent
- Is used to explain the fact that the carbon-to-carbon bonds in benzene, C_6H_6 , are identical

Directions: Each of the questions or incomplete statements below is followed by five suggested answers or compilations. Select the one that is best in each case and then blacken the corresponding space on the answer sheet.

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Howard Wainer



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1. Understanding the eBook 1989 Ap Chemistry Multiple Choice
 - The Rise of Digital Reading 1989 Ap Chemistry Multiple Choice
 - Advantages of eBooks Over Traditional Books
2. Identifying 1989 Ap Chemistry Multiple Choice
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 1989 Ap Chemistry Multiple Choice
 - User-Friendly Interface
4. Exploring eBook Recommendations from 1989 Ap Chemistry Multiple Choice
 - Personalized Recommendations
 - 1989 Ap Chemistry Multiple Choice User Reviews and Ratings

- 1989 Ap Chemistry Multiple Choice and Bestseller Lists
- 5. Accessing 1989 Ap Chemistry Multiple Choice Free and Paid eBooks
 - 1989 Ap Chemistry Multiple Choice Public Domain eBooks
 - 1989 Ap Chemistry Multiple Choice eBook Subscription Services
 - 1989 Ap Chemistry Multiple Choice Budget-Friendly Options
- 6. Navigating 1989 Ap Chemistry Multiple Choice eBook Formats
 - ePub, PDF, MOBI, and More
 - 1989 Ap Chemistry Multiple Choice Compatibility with Devices
 - 1989 Ap Chemistry Multiple Choice Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 1989 Ap Chemistry Multiple Choice
 - Highlighting and Note-Taking 1989 Ap Chemistry Multiple Choice
 - Interactive Elements 1989 Ap Chemistry Multiple Choice
- 8. Staying Engaged with 1989 Ap Chemistry Multiple Choice
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 1989 Ap Chemistry Multiple Choice
- 9. Balancing eBooks and Physical Books 1989 Ap Chemistry Multiple Choice
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 1989 Ap Chemistry Multiple Choice
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 1989 Ap Chemistry Multiple Choice
 - Setting Reading Goals 1989 Ap Chemistry Multiple Choice
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 1989 Ap Chemistry Multiple Choice
 - Fact-Checking eBook Content of 1989 Ap Chemistry Multiple Choice
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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