



General Chemistry II Equation Sheet

Think about how to set up the problem first, then apply the needed principles and formulas.

Phase Changes

$$\Delta T_b = iK_b m$$

$$\Delta T_f = iK_f m$$

$$\pi = iMRT$$

$$c = kP$$

$$\frac{P_1}{P_2} = \frac{r_1}{r_2}$$

$$P_{\text{solution}} = \chi_{\text{solvent}} P_{\text{solvent}}^{\circ}$$

$$\Delta P = P_{\text{solvent}}^{\circ} - P_{\text{solution}} = \chi_{\text{solute}} P_{\text{solvent}}^{\circ}$$

Cells

$$E = E^{\circ} - \frac{RT}{nF} \ln(Q)$$

$$E = E^{\circ} - \frac{0.0592}{n} \log(Q) \quad (\text{at } 25^{\circ}\text{C})$$

$$\Delta G = -nFE_{\text{cell}}$$

$$\Delta G^{\circ} = -nFE_{\text{cell}}^{\circ}$$

$$E_{\text{cell}}^{\circ} = \frac{RT}{nF} \ln(K)$$

$$E_{\text{cell}}^{\circ} = \frac{0.0592\text{V}}{n} \log(K) \quad (\text{at } 25^{\circ}\text{C})$$

USE ONE:

$$E_{\text{cell}}^{\circ} = E_{\text{cathode(re.d.)}}^{\circ} - E_{\text{anode(re.d.)}}^{\circ}$$

$$E_{\text{cell}}^{\circ} = E_{\text{cathode(re.d.)}}^{\circ} + E_{\text{anode(ox.)}}^{\circ}$$

Thermodynamics

$$\Delta G = \Delta H - T\Delta S$$

$$\Delta S_{\text{surr}} = \frac{-\Delta H_{\text{sys}}}{T}$$

$$\Delta S_{\text{univ}} = \Delta S_{\text{sys}} + \Delta S_{\text{surr}}$$

$$\Delta S_{\text{ran}}^{\circ} = \Sigma S_{\text{products}}^{\circ} - \Sigma S_{\text{reactants}}^{\circ}$$

$$\Delta G_{\text{ran}}^{\circ} = \Sigma \Delta G_{\text{products}}^{\circ} - \Sigma \Delta G_{\text{reactants}}^{\circ}$$

Concentrations

$$\text{molality} = \frac{\text{moles solute}}{\text{mass solvent (kg)}}$$

$$\% \text{ by mass} = \frac{\text{mass solute}}{\text{mass solute} + \text{mass solvent}} \times 100\%$$

$$\text{parts per million} = \frac{\text{mass solute}}{\text{mass solute} + \text{mass solvent}} \times 10^6$$

$$\text{parts per billion} = \frac{\text{mass solute}}{\text{mass solute} + \text{mass solvent}} \times 10^9$$

Acids and Bases

$$K_w = [H_3O^+][OH^-] = K_a \times K_b$$

$$pOH = -\log[OH^-]$$

$$[OH^-] = 10^{-pOH}$$

$$pH + pOH = pK_a + pK_b = 14.00 \quad (\text{at } 25^{\circ}\text{C})$$

$$pH = pK_a + \log \frac{[A^-]}{[HA]}$$

$$pOH = pK_b + \log \frac{[BH^+]}{[B]}$$

$$\% \text{ ionization} = \frac{[H_3O^+]_{\text{eq}}}{[HA]_{\text{in}}} \times 100\%$$

$$pK_a = -\log(K_a)$$

$$pK_b = -\log(K_b)$$

$$K_a = 10^{-pK_a}$$

$$K_b = 10^{-pK_b}$$

Equilibria

$$Q_c = \frac{[C]^c[D]^d}{[A]^a[B]^b}$$

$$K_c = \frac{[C]^c[D]^d}{[A]^a[B]^b}$$

$$Q_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$$

$$K_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$$

$$K_p = K_c [R \times T]^{\Delta n}$$

$$\Delta G = \Delta G^{\circ} + RT \ln(Q)$$

$$\Delta G^{\circ} = -RT \ln(K)$$

$$\text{rate} = k[A]^x[B]^y$$

$$k = Ae^{-\frac{E_a}{RT}}$$

$$\ln\left(\frac{k_2}{k_1}\right) = \frac{E_a}{R} \left(\frac{1}{T_1} - \frac{1}{T_2}\right)$$

$$\text{half life} = \frac{\ln(2)}{k} \quad (\text{1st order reactions only})$$

$$[A]_t = -kt + [A]_0$$

$$\ln([A]_t) = -kt + \ln([A]_0)$$

$$\frac{1}{[A]_t} = kt + \frac{1}{[A]_0}$$

Answer Sheet General Chemistry Equations

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Answer Sheet General Chemistry Equations:

General Chemistry Workbook Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material *Chemical Symbolism and Calculations* Stanley Wallace Morse, 1953 *Laboratory Manual for Principles of General Chemistry* Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures **Laboratory Manual for Principles of General Chemistry** J. A. Beran, Mark Lassiter, 2022-08-16 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition's experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section **General Chemistry** Donald A. McQuarrie, Stanley Gill, 2011-06-15 This Fourth Edition of McQuarrie's classic text offers a thorough revision and a quantum leap forward from the previous edition Taking an atoms first approach it promises to be another ground breaking text in the tradition of McQuarrie's many previous works This outstanding new text available in a soft cover edition offers professors a fresh choice and outstanding value *Regents Exams and Answers: Chemistry--Physical Setting Revised Edition* Barron's Educational Series, Albert Tarendash, 2021-01-05 Barron's Regents Exams and Answers Chemistry provides essential practice for students taking the

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Strategies and Solutions to Advanced Organic Reaction Mechanisms Andrei

Hent, John Andraos, 2019-06-26 *Strategies and Solutions to Advanced Organic Reaction Mechanisms A New Perspective on McKillop's Problems* builds upon Alexander Sandy McKillop's popular text *Solutions to McKillop's Advanced Problems in Organic Reaction Mechanisms* providing a unified methodological approach to dealing with problems of organic reaction mechanism. This unique book outlines the logic, experimental insight and problem solving strategy approaches available when dealing with problems of organic reaction mechanism. These valuable methods emphasize a structured and widely applicable approach relevant for both students and experts in the field. By using the methods described, advanced students and researchers alike will be able to tackle problems in organic reaction mechanism from the simple and straight forward to the advanced.

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CBSE - ICSE Chemistry Part I Chandan Sengupta, First Publication October 2021 Place of Publication Arabinda Nagar Bankura 722101 This workbook will provide an ample scope in getting exposed to the system of acquiring skills and competence related to the understanding of chemistry. It also exposes the student to the concepts of chemistry for enabling the aspirant in acquisition of skills related to chemistry. Some of the worksheets are prepared along with model answers. Some other worksheets are meant for self assessment and evaluation purposes. It is also observed that some of the topics are specific to the referred curriculum. Some other topics are varyingly incorporated in other streams of study. Culmination of more than two streams will enable the fellow student to cope up with the preparatory works meant for Olympiads and other competitive examinations.

A Laboratory Outline of General Chemistry Alexander Smith, 1899

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