

OpenQuest 13

Orbital Configuration

Name _____

Date _____

Class _____

Information: Energy of Sublevels

Not all orbitals have the same energy of energy. For example, orbitals in the $3p$ sublevel have more energy than orbitals in the $3s$ sublevel. The following orbitals have different energy levels in highest energy.



Circle the orbitals that have the same energy level. (Remember, orbitals in the same sublevel have the same energy.)



Now list the orbitals that have the same energy level. (Remember, orbitals in the same sublevel have the same energy.)

When electrons occupy orbitals, they try to have the lowest possible energy possible. (This is called the **lowest energy principle**.) An electron will never be added to an orbital until the orbital is full. (This is called the **Pauli Exclusion Principle**.) When two electrons occupy an orbital, they must spin in opposite directions—orbital spins must be opposite.

Critical Thinking Questions

- a. How many electrons could a sublevel hold if it were full? (Remember, orbitals in the same sublevel have the same energy.)
b. How many electrons could a sublevel hold if it were full? (Remember, orbitals in the same sublevel have the same energy.)
- a. How many electrons could a sublevel hold if it were full? (Remember, orbitals in the same sublevel have the same energy.)
b. How many electrons could a sublevel hold if it were full? (Remember, orbitals in the same sublevel have the same energy.)

Information: Hund's Rule

Electrons prefer to "spin" or "upspin". (Remember, orbitals in the same sublevel have the same energy.)



Remember, electrons in the same sublevel have the same energy. (Remember, orbitals in the same sublevel have the same energy.)



Hund's rule says that electrons in the same sublevel will have the same energy. (Remember, orbitals in the same sublevel have the same energy.)

The following are the electron orbital diagrams for the elements, nitrogen and oxygen. (Remember, orbitals in the same sublevel have the same energy.)



Critical Thinking Questions

- a. How many "upspin" electrons are in a sublevel? (Remember, orbitals in the same sublevel have the same energy.)
- a. Why does oxygen's electron distribution look like this? (Remember, orbitals in the same sublevel have the same energy.)
- a. Why does nitrogen's electron distribution look like this? (Remember, orbitals in the same sublevel have the same energy.)
- a. Why does oxygen's electron distribution look like this? (Remember, orbitals in the same sublevel have the same energy.)

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AN Whitehead



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