

Gravimetric Analysis of Chloride in Solution

Experiment 10
Lab 10
Date: _____

Introduction

This experiment involves the gravimetric analysis of a solution containing chloride ions. The procedure involves the precipitation of the chloride ions as silver chloride (AgCl) using a solution of silver nitrate (AgNO_3). The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution. The experiment is performed in two parts. In the first part, a known mass of a sample containing chloride ions is dissolved in water. In the second part, a known volume of a silver nitrate solution is added to the sample solution, causing the precipitation of silver chloride. The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution.

The experiment is performed in two parts. In the first part, a known mass of a sample containing chloride ions is dissolved in water. In the second part, a known volume of a silver nitrate solution is added to the sample solution, causing the precipitation of silver chloride. The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution.

Procedure

The procedure involves the gravimetric analysis of a solution containing chloride ions. The procedure involves the precipitation of the chloride ions as silver chloride (AgCl) using a solution of silver nitrate (AgNO_3). The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution.

Data Analysis

	Sample 1	Sample 2
Mass of sample	0.100 g	0.100 g
Mass of precipitate	0.143 g	0.143 g
Mass of chloride ion	0.050 g	0.050 g

The mass of chloride ion in the sample is calculated as follows:

$$\text{Mass of Cl}^- = \text{Mass of precipitate} \times \frac{\text{Molar mass of Cl}^-}{\text{Molar mass of AgCl}}$$

The mass of chloride ion in the sample is calculated as follows:

Analysis Of Chloride Lab Report

Guillaume Favre



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