# **Balancing redox equations**

## Example

Balance the equation showing the oxidation of Fe<sup>2+</sup> ions to Fe<sup>3+</sup> ions by dichromate ions (Cr<sub>2</sub>O<sub>7</sub><sup>2-</sup>) in an acidic medium

$$Fe^{2+} + Cr_2O_7^{2-} \rightarrow Fe^{3+} + Cr^{3+}$$

Step 1: Identify oxidising and reducing agents and write half reactions

Fe<sup>2+</sup>: +2 
$$Cr_2O_7^{2-}$$
:  $2Cr + 7(-2) = -2$   
Fe<sup>3+</sup>: +3  $2Cr = 12$   
 $Cr^{3+}$ : +3  $Cr = +6$ 

$$Fe^{2+} \rightarrow Fe^{3+}$$

Iron loses one electron, i.e. it is oxidised. It donates one electron to dichromate, i.e. it is the reducing agent.

$$Cr_2^{6+} \rightarrow Cr^{3+}$$

Chromium gains three electrons, i.e. it is reduced. It gains its electrons from iron, i.e. it is the oxidising agent.

Half equation: Reduction reaction: Cr26+ + 3e- → Cr3+

## **Analyzing Oxidation Reduction Reactions Answers**

**Julius Stieglitz** 

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