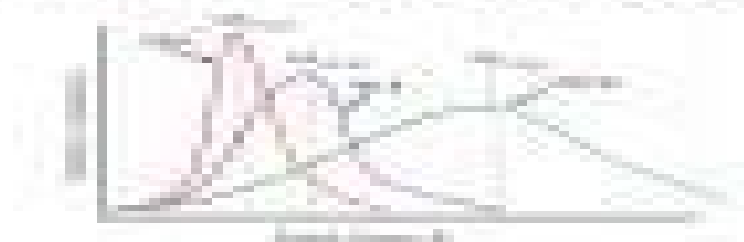


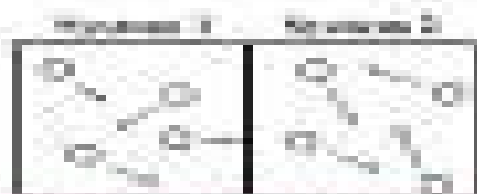
Unit 6: Thermochemistry**6.1: Heat Transfer and Thermal Equilibrium****6.2: Heat Capacity and Calorimetry****6.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving with v speed is $\frac{1}{2}mv^2$, the first being the mass and momentum being the change, directly connected to all of temperature increase. The Boltzmann distribution curve shows that the average velocity is small for the distribution's maximum value to change the average temperature increases.

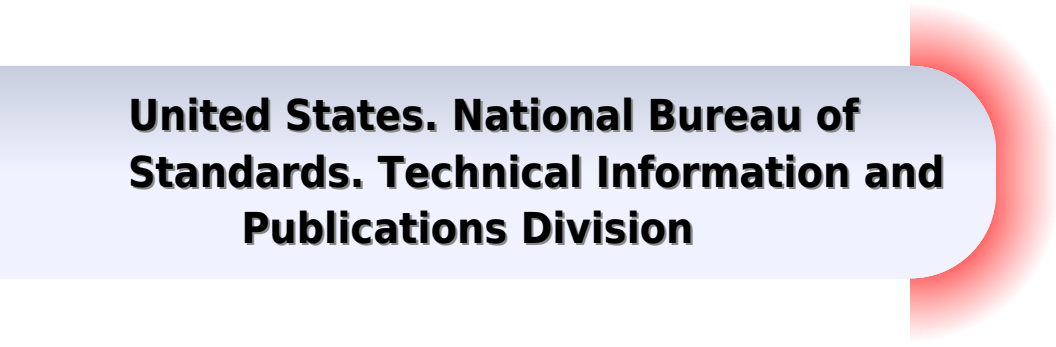
3. Sketch a curve used to represent the relative velocities of the particles in the following two gas samples of the same gas.



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.
 System 1 is at a lower temperature than System 2. Temperature is a measure of the average KE of the particles in the system. In System 1, the particles have low kinetic energy, and all the individual particles have low kinetic energy. According to the equation $KE = \frac{1}{2}mv^2$, all of the particles have the same mass, so the only way to get a low KE is for the velocity to be low. In System 2, the particles are moving at much higher speeds, so the average KE is higher.
- b. Are the particles in both boxes moving faster or slower than? Justify your answer in terms of molecular collisions.
 All particles in both boxes are moving faster than System 1. In System 1, the particles have low kinetic energy, and all the individual particles have low kinetic energy. According to the equation $KE = \frac{1}{2}mv^2$, all of the particles have the same mass, so the only way to get a low KE is for the velocity to be low. In System 2, the particles are moving at much higher speeds, so the average KE is higher.

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