

Lab 10: Calculations with a Chemical Structure

Objectives:

1. To derive an empirical formula from a chemical structure without performing stoichiometric calculations.
2. To calculate the number of moles of a solid from a weighing measurement in the lab.
3. To determine the molar mass of a solid.
4. To determine the theoretical amount of product that would be produced.
5. To compare the theoretical amount to the actual amount of substance formed and calculate the percent yield.

Reaction equation:

Reaction of sodium chloride solution: 25.0 mL

Reaction of calcium chloride solution: 25.0 mL

Reaction of silver nitrate: 5.00 g

Reaction of silver nitrate: 5.00 g

Reaction of silver nitrate: 5.00 g

Reaction of silver nitrate: 5.00 g

Procedure:

Reaction and Calculations:

1. Reaction:

$$\text{mole of sodium chloride} = 25.0 \text{ mL sol. } \times \frac{1.0}{1000 \text{ mL}} = 0.025 \text{ mole NaCl}$$

$$\text{mole of calcium chloride} = 25.0 \text{ mL sol. } \times \frac{1.0}{1000 \text{ mL}} = 0.025 \text{ mole CaCl}_2$$

$$\text{mole of silver nitrate produced} = 5.00 \text{ g} \times \frac{1 \text{ mole}}{169.87 \text{ g}} = 0.0295 \text{ mole AgNO}_3$$

$$\text{mole of product that would be produced} = 0.025 \text{ mole CaCl}_2 \times \frac{1 \text{ mole}}{1 \text{ mole}} = 0.025 \text{ mole CaCl}_2$$

2. Reaction equation: $\text{CaCl}_2 + 2\text{AgNO}_3 \rightarrow \text{Ca(NO}_3)_2 + 2\text{AgCl}$

3. Reaction equation:

$$\text{mole of } \text{CaCl}_2 = 0.025 \text{ mole} \times \frac{1 \text{ mole}}{1 \text{ mole}} = 0.025 \text{ mole CaCl}_2$$

$$\text{mole of } \text{CaCl}_2 = 0.025 \text{ mole}$$

$$\text{mole of } \text{CaCl}_2 = 0.025 \text{ mole} \times \frac{1 \text{ mole}}{1 \text{ mole}} = 0.025 \text{ mole CaCl}_2$$

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