

CHAPTER 10: Reactions

Reaction Energy**SECTION 1**

Write a balanced chemical equation for the following reactions in the space provided.

- For elements in their standard state, the value of ΔH_f° is _____.
- The formation and decomposition of a substance is represented by the following thermodynamic equations:



- Is the energy being added to or is it being subtracted from the decomposition?
- What is the correct sign for the enthalpy change in this decomposition reaction?

Write the correct answer in the space to the left. Show all your work in the space provided.

- _____ $\Delta H_{\text{f}}^\circ$ of a substance is 2074°C whereas a 4.000 g of substance reacts with O_2 to form CO_2 and H_2O . What is the $\Delta H_{\text{f}}^\circ$ of the substance?
- _____ A substance has a specific heat of $0.001 \text{ J/g}^\circ\text{C}$. If a 1.00 g sample of substance is heated from 25°C to 100°C , how much heat is absorbed? (The substance has a $\Delta H_{\text{f}}^\circ$ of 100 J/mol .)
- The product is a molecular compound of 10.0 g/mol, and the substance has a molar mass of 100.0 g/mol.
 - What is the value of $\Delta H_{\text{f}}^\circ$ for this reaction?

Answers For Reaction Energy Section 1

M Woodhall



Answers For Reaction Energy Section 1:

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