

CHEM 151 – Quiz #4
Thermo, Redox, and Electrochem
 December 4, 2018

1. Consider this unbalanced reaction at 25 °C:

Zinc solid reacts with a solution containing nitrate ion, to form zinc ion and nitrogen monoxide gas.

The standard reduction potentials (E°) are:

Nitrogen half-reaction:	+ 0.957 V	red
Zinc half-reaction:	- 0.762 V	ox

- a. (2 pts) Using chemical formulas, write the unbalanced reaction given above.



- b. (5 pts) Balance the reaction in an acidic solution using the half-reaction method practiced in class (include phases in your final balanced reaction).



- c. (3 pts) i) assign oxidation numbers to all atoms in your final balanced reaction, ii) calculate the change of oxidation state for the two atoms that underwent change, and identify if each atom underwent oxidation or reduction, and iii) label the oxidizing and reducing agents (from the reactants only).

i) see above

ii) $\text{Zn} \xrightarrow{\text{iii)}} 0 \rightarrow +2$, oxidation $\Rightarrow \text{Zn(s)} = \text{reducing agent}$

$\text{N} \xrightarrow{\text{iii)}} +5 \rightarrow +2$, reduction $\Rightarrow \text{NO}_3^-(\text{aq}) = \text{oxidizing agent}$

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Table of Contents Answers For Chemistry 151 Lab

1. Understanding the eBook Answers For Chemistry 151 Lab
 - The Rise of Digital Reading Answers For Chemistry 151 Lab
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers For Chemistry 151 Lab
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers For Chemistry 151 Lab
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers For Chemistry 151 Lab
 - Personalized Recommendations
 - Answers For Chemistry 151 Lab User Reviews and Ratings
 - Answers For Chemistry 151 Lab and Bestseller Lists
5. Accessing Answers For Chemistry 151 Lab Free and Paid eBooks
 - Answers For Chemistry 151 Lab Public Domain eBooks
 - Answers For Chemistry 151 Lab eBook Subscription Services
 - Answers For Chemistry 151 Lab Budget-Friendly Options

6. Navigating Answers For Chemistry 151 Lab eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers For Chemistry 151 Lab Compatibility with Devices
 - Answers For Chemistry 151 Lab Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers For Chemistry 151 Lab
 - Highlighting and Note-Taking Answers For Chemistry 151 Lab
 - Interactive Elements Answers For Chemistry 151 Lab
8. Staying Engaged with Answers For Chemistry 151 Lab
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers For Chemistry 151 Lab
9. Balancing eBooks and Physical Books Answers For Chemistry 151 Lab
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers For Chemistry 151 Lab
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answers For Chemistry 151 Lab
 - Setting Reading Goals Answers For Chemistry 151 Lab
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answers For Chemistry 151 Lab
 - Fact-Checking eBook Content of Answers For Chemistry 151 Lab
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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