

# Stoichiometry Limiting Reactant

## Answers

- (1) How many grams of calcium phosphate can be produced from the reaction between 2.50 L of 0.250 M calcium chloride with an excess of phosphoric acid?



$$2.50 \text{ L CaCl}_2 \times \frac{0.250 \text{ mol CaCl}_2}{1 \text{ L CaCl}_2} \times \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{3 \text{ mol CaCl}_2} \times \frac{310 \text{ g Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} = 64.6 \text{ g Ca}_3(\text{PO}_4)_2$$

- (2) How many milliliters of 1.50 M nitric acid is required to react with 100.0 g of cuprous oxide?



$$100 \text{ g Cu}_2\text{O} \times \frac{1 \text{ mol Cu}_2\text{O}}{143 \text{ g Cu}_2\text{O}} \times \frac{14 \text{ mol HNO}_3}{3 \text{ mol Cu}_2\text{O}} \times \frac{1000 \text{ ml HNO}_3}{1.5 \text{ mol HNO}_3} = 2180 \text{ ml HNO}_3$$

- (3) Consider the following unbalanced reaction:



Balance the reaction. 3.25 g of  $\text{NH}_3$  is allowed to react with 3.5 g of  $\text{O}_2$ .

- (a) Which reactant is the limiting reagent?

$$3.25 \text{ g NH}_3 \times \frac{1 \text{ mol NH}_3}{17.03 \text{ g NH}_3} \times \frac{4 \text{ mol NO}}{4 \text{ mol NH}_3} \times \frac{30.01 \text{ g NO}}{1 \text{ mol NO}} = 5.73 \text{ g NO}$$

$$3.5 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{4 \text{ mol NO}}{5 \text{ mol O}_2} \times \frac{30.01 \text{ g NO}}{1 \text{ mol NO}} = 2.63 \text{ g NO}$$

$\text{O}_2$  is the limiting reagent.

- (b) How many grams of NO are formed?

2.63 g of NO are formed.

- (c) How much excess reagent remains after the reaction?

$$3.5 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{4 \text{ mol NH}_3}{5 \text{ mol O}_2} \times \frac{17.03 \text{ g NH}_3}{1 \text{ mol NH}_3} = 1.49 \text{ g NH}_3 \text{ used}$$

$$3.25 \text{ g NH}_3 - 1.49 \text{ g NH}_3 = 1.76 \text{ g NH}_3 \text{ remains.}$$

# Answer Sheet To The Stoichiometry Limiting Reagent

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Chemistry John Olmsted,Greg Williams,Robert C. Burk,2020 Chemistry 4th Edition is an introductory general chemistry text designed specifically with Canadian professors and students in mind A reorganized Table of Contents and inclusion of SI units IUPAC standards and Canadian content designed to engage and motivate readers and distinguish this text from other offerings It more accurately reflects the curriculum of most Canadian institutions Chemistry is sufficiently rigorous while engaging and retaining student interest through its accessible language and clear problem solving program without an excess of material and redundancy

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Arihant Experts, 2019-06-04 While beginning the preparation for Medical and Engineering Entrances aspirants need to go beyond traditional NCERT textbooks to gain a complete grip over it to answer all questions correctly during the exam The revised edition of MASTER THE NCERT based on NCERT Classes XI and XII once again brings a unique set of all kinds of Objective Type Questions for Physics Chemistry Biology and Mathematics This book Master the NCERT for NEET Chemistry Vol 1 based on NCERT Class XI is a one of its kind book providing 14 Chapters equipped with topic wise objective questions NCERT Exemplar Objective Questions and a special separate format questions for NEET and other medical entrances It also provides explanations for difficult questions and past exam questions for knowing the pattern Based on a unique approach to master NCERT it is a perfect study resource to build the foundation over NEET and other medical entrances

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John Christian Bailar, 1984

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