

SCHAU Unit 4 - Quiz – Equilibrium

Player(s): _____

Period(s): _____

Q1: (10)

Q2: (10)

Q3: (10)

Q4: (10)

Q5: (10)

Relevant Equations:

$$K_c = \frac{[\text{products}]}{[\text{reactants}]} \quad K_p = \frac{(P)^{\text{products}}}{(P)^{\text{reactants}}}$$

$$K_c = \frac{P_{\text{CO}_2}^2 P_{\text{H}_2\text{O}}}{P_{\text{CH}_4}^2 P_{\text{H}_2\text{O}}^2} \quad K_p = \frac{(P_{\text{CO}_2})^2 (P_{\text{H}_2\text{O}})}{(P_{\text{CH}_4})^2 (P_{\text{H}_2\text{O}})^2}$$

gib (1 point) : 100 : gib (10) : gib (10) : gib (10) : gib (10) : gib (10)

100 : 100 : 100

Question 1

1. Hydrogen sulfide, H_2S , is a slightly poisonous gas. At 1000 K, an equilibrium mixture contains 0.020 moles of hydrogen, H_2 , 0.010 moles of hydrogen sulfide, with an equilibrium pressure of 0.020 atm. The reaction is:



If the volume of an 8.00 L flask at 1000 K where a concentration of H_2 is present at equilibrium:

- What is the concentration of H_2 at equilibrium?
 Answer: $[\text{H}_2] = 0.0025 \text{ M}$
 Answer: $[\text{H}_2] = 0.0025 \text{ M}$
 $[\text{H}_2] = 0.0025 \text{ M}$
- Calculate the value of K_c at equilibrium.
- Answer:
 What other species have chemical formula and a 2 in front coefficient above?
- Calculate:

	$[\text{H}_2\text{S}]_{\text{eq}}$	$[\text{H}_2]_{\text{eq}}$	$[\text{S}_2]_{\text{eq}}$
Initial (moles/liter)			
Change (moles/liter)			
Eq. (moles/liter)	0.010	0.020	0.005

1000 K, calculate:

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