

① $x = 0.01$ Cr, $Ni = 0.04$

Mg = $0.5 \times 0.95 = 0.475$ M (Mg nitrate)

$0.475 = \frac{n}{0.05} \Rightarrow n = 0.02375$ mol

$\therefore w = 6.0895$ g
Mg nitrate

Ni = $0.5 \times 0.04 = 0.02$ M

$0.02 = \frac{n}{0.05} \Rightarrow n = 0.001$ mol

$w = 0.2908$ g
Ni nitrate

Cr = $0.5 \times 0.01 = 0.005$

$0.005 = \frac{n}{0.05} \Rightarrow n = 0.00025$ mol

$w = 0.10035$ g
Cr nitrate

② $x = 0.03$ Cr, $Ni = 0.02$

Mg $w = 6.0895$ g

Ni = $0.5 \times 0.02 = 0.01$ M

$\therefore n = 0.0005$ mol Ni

$w = 0.1457$ g
Ni nitrate

Cr = $0.5 \times 0.03 = 0.015$ M

$n = 0.00075$ mol

$w = 0.300105$ g
Cr nitrate

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