

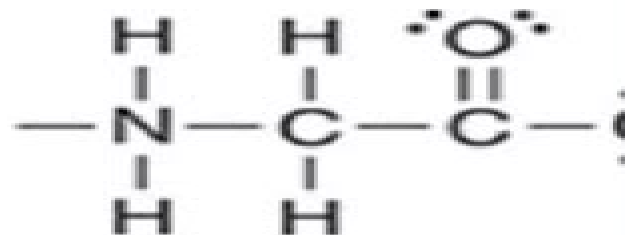
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Which is the correct hybridization state and geometry for the carbon atom in HCN?

- sp, linear
- sp², trigonal planar
- sp³, tetrahedral
- None of the above
- sp, linear

The following structure has been drawn without formal charges. Which statement describes the missing formal charge(s)?

- This structure has one positive charge and one negative charge.
- This structure has one positive charge but no negative charges.
- This structure has one negative charge but no positive charges.
- This structure has no formal charges.
- This structure has one positive charge and one negative charge.



The indicated σ bond results from the overlap of which orbitals?

- sp²-sp²
- sp-sp³
- sp-sp²
- sp²-sp³
- sp²-sp³

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