

More aqueous equilibrium self test/practice problems, Dr. Scott Ensign.

1. **(Acid-base equilibrium review)** A solution is prepared by dissolving 0.15 moles of acetic acid in 1.00 liters of solution.  $K_a$  for acetic acid is  $1.8 \times 10^{-5}$ . What will be the pH of the solution once equilibrium is established?

- a. 2.0
- b. 5.6
- c. 2.8
- d. 3.9
- e. 4.7

2. **(Acid-base equilibrium review)** A solution is prepared by dissolving 0.10 moles of sodium acetate ( $\text{NaCH}_3\text{COO}$ ) in 1.00 liters of solution.  $K_a$  for acetic acid is  $1.8 \times 10^{-5}$ . What will be the pH of the solution once equilibrium is established?

- a. 8.9
- b. 5.1
- c. 9.9
- d. 4.1
- e. 9.3

3. A solution is prepared by mixing 0.15 moles of acetic acid ( $\text{CH}_3\text{COOH}$ ) with 0.10 moles of sodium acetate ( $\text{NaCH}_3\text{COO}$ ) in 1.00 liters of solution.  $K_a$  for acetic acid is  $1.8 \times 10^{-5}$ . What will be the pH of the solution once equilibrium is established?

- a. 4.9
- b. 6.6
- c. 2.9
- d. 4.6
- e. 4.7

4. A solution is prepared by mixing 0.15 moles of acetic acid ( $\text{CH}_3\text{COOH}$ ) with 0.050 moles of nitric acid ( $\text{HNO}_3$ ) in 1.00 liters of solution.  $K_a$  for acetic acid is  $1.8 \times 10^{-5}$ . What will be the pH of the solution once equilibrium is established?

- a. 2.4
- b. 2.8
- c. 1.3
- d. 4.3
- e. 5.2

5. **(Acid-base equilibrium review)** What would be the pH of a solution prepared by dissolving 0.55 moles of  $\text{NH}_3$  in 1.00 liters of solution?  $K_b$  for  $\text{NH}_3$  is  $1.80 \times 10^{-5}$

- a. 9.51
- b. 11.5
- c. 11.8
- d. 2.50
- e. 9.00

# Aqueous Equilibrium Practice Problems

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Williams, Robert Charles Burk**



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