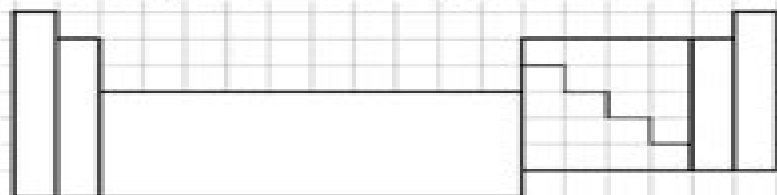


Exam #3 Review:

Periodic Table Basics, Electron Configuration, and Periodic Trends

Periodic Table Basics:

- How did Mendeleev arrange the elements?
How did Moseley organize the periodic table?
Which way is it organized today?
- What is the difference between groups and periods on the periodic table?
- Label the 5 major families on the periodic table:



- What side of the stair-step are the metals found on?
What side of the stair-step are the nonmetals found on?
- Where are the metalloids located? What 7 elements are metalloids?
- Why do the elements potassium and sodium have similar physical and chemical properties?
- How do you know how many **valence** electrons are in an atom? (What do you have to look at on the periodic table?)

Electron Configuration:

- What is the name of the modern atomic model?
- Know the 3 rules that electrons follow when filling up sublevels:

Aufbau Principle –

Hund's Rule –

Pauli Exclusion Principle –

- Fill in the following table with the information for each sublevel:

Sublevel	Shape	Maximum # of Electrons	# of Orbitals
s			
p			
d			
f			

- Which sublevels are found in the following?

- 1st energy level:
- 2nd energy level:
- 3rd energy level:

Are You Able With A Table Answer Key

Robert L. Schalock



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