

OpenQuest 13

Orbital Configuration

Name _____

Date _____

Class _____

Information: Energy of Sublevels

Sub-levels have different amounts of energy. For example, orbitals closer to the nucleus have more energy than orbitals further from the nucleus. The following orbitals have different levels of energy:



For each pair, write sublevel(s) for which orbitals are included. Then draw orbital(s) representing an orbital.



Now list 4 orbitals which appear in the set of pairs. Write down their approximate energy levels. For example, the 3d orbitals have higher energy than s orbitals and the 4f orbitals have a higher energy than d orbitals.

When electrons occupy orbitals, they try to have the lowest amount of energy possible. (This is called the **lowest-energy principle**.) An electron will occupy 1 orbital before the 1 orbital is filled and no electron will occupy an orbital until it is completely filled. Additionally, no electron will be sub-filled. (This is called the **Mutual Exclusion Principle**.) When two electrons occupy the same orbital they must spin in opposite directions – one clockwise and the other anticlockwise.

Critical Thinking Questions

- a) How many electrons would it take to fill a sublevel before beginning to be sub-filled? **Sublevels become filled if the orbitals are filled before electrons reach orbitals in the 2p orbitals.**
b) What is the first element before enough electrons are present to fill 1 orbital? **After you provide this:**
The 1st element before this is 1 orbital. The element is 1 electron 1 orbital.
- a) How many electrons would it take to fill a sublevel before beginning to be sub-filled? **Sublevels become filled if the orbitals are filled before electrons reach orbitals in the 2p orbitals.**
b) What is the first element before enough electrons are beginning to fill orbitals in the 2p orbitals? **The 2nd element before this is 2 orbitals. The element is 2 electrons 2 orbitals.**

Information: Hund's Rule

Electrons prefer to "spin" or "upspin". Orbitals closer to the nucleus have a higher potential. Electrons closer to the nucleus have a lower potential. Electrons closer to the nucleus have a lower potential.



If an orbital is represented as a box, the orbital diagram will get electron's orbital diagram. The diagram represents a box for each orbital. The diagram represents a box for each orbital.



Hund's rule tells us that orbitals closer to the nucleus have a higher potential. Orbitals closer to the nucleus have a higher potential. Orbitals closer to the nucleus have a higher potential. Orbitals closer to the nucleus have a higher potential.

The following are the electron orbital diagrams for the elements, nitrogen and oxygen. Write the diagram's 2p sublevel and explain to show Hund's rule. The 2p sublevel is filled before beginning to fill orbitals in the 2p orbitals.



Critical Thinking Questions

- How many "upspin" electrons are in a nitrogen atom? **1**
- Why does carbon's orbitals diagram is given as 1s² 2s² 2p²? Why not 1s² 2s² 2p²? **The 2p orbitals are already full. The 2p orbitals are already full. The 2p orbitals are already full.**
- Why do elements before nitrogen have 1s² 2s² 2p³ 3s² 3p² 3d² 4s² 4p² 4d² 4f² 5s² 5p² 5d² 5f² 6s² 6p² 6d² 6f² 7s² 7p² 7d² 7f² 8s² 8p² 8d² 8f² 9s² 9p² 9d² 9f² 10s² 10p² 10d² 10f² 11s² 11p² 11d² 11f² 12s² 12p² 12d² 12f² 13s² 13p² 13d² 13f² 14s² 14p² 14d² 14f² 15s² 15p² 15d² 15f² 16s² 16p² 16d² 16f² 17s² 17p² 17d² 17f² 18s² 18p² 18d² 18f² 19s² 19p² 19d² 19f² 20s² 20p² 20d² 20f² 21s² 21p² 21d² 21f² 22s² 22p² 22d² 22f² 23s² 23p² 23d² 23f² 24s² 24p² 24d² 24f² 25s² 25p² 25d² 25f² 26s² 26p² 26d² 26f² 27s² 27p² 27d² 27f² 28s² 28p² 28d² 28f² 29s² 29p² 29d² 29f² 30s² 30p² 30d² 30f² 31s² 31p² 31d² 31f² 32s² 32p² 32d² 32f² 33s² 33p² 33d² 33f² 34s² 34p² 34d² 34f² 35s² 35p² 35d² 35f² 36s² 36p² 36d² 36f² 37s² 37p² 37d² 37f² 38s² 38p² 38d² 38f² 39s² 39p² 39d² 39f² 40s² 40p² 40d² 40f² 41s² 41p² 41d² 41f² 42s² 42p² 42d² 42f² 43s² 43p² 43d² 43f² 44s² 44p² 44d² 44f² 45s² 45p² 45d² 45f² 46s² 46p² 46d² 46f² 47s² 47p² 47d² 47f² 48s² 48p² 48d² 48f² 49s² 49p² 49d² 49f² 50s² 50p² 50d² 50f² 51s² 51p² 51d² 51f² 52s² 52p² 52d² 52f² 53s² 53p² 53d² 53f² 54s² 54p² 54d² 54f² 55s² 55p² 55d² 55f² 56s² 56p² 56d² 56f² 57s² 57p² 57d² 57f² 58s² 58p² 58d² 58f² 59s² 59p² 59d² 59f² 60s² 60p² 60d² 60f² 61s² 61p² 61d² 61f² 62s² 62p² 62d² 62f² 63s² 63p² 63d² 63f² 64s² 64p² 64d² 64f² 65s² 65p² 65d² 65f² 66s² 66p² 66d² 66f² 67s² 67p² 67d² 67f² 68s² 68p² 68d² 68f² 69s² 69p² 69d² 69f² 70s² 70p² 70d² 70f² 71s² 71p² 71d² 71f² 72s² 72p² 72d² 72f² 73s² 73p² 73d² 73f² 74s² 74p² 74d² 74f² 75s² 75p² 75d² 75f² 76s² 76p² 76d² 76f² 77s² 77p² 77d² 77f² 78s² 78p² 78d² 78f² 79s² 79p² 79d² 79f² 80s² 80p² 80d² 80f² 81s² 81p² 81d² 81f² 82s² 82p² 82d² 82f² 83s² 83p² 83d² 83f² 84s² 84p² 84d² 84f² 85s² 85p² 85d² 85f² 86s² 86p² 86d² 86f² 87s² 87p² 87d² 87f² 88s² 88p² 88d² 88f² 89s² 89p² 89d² 89f² 90s² 90p² 90d² 90f² 91s² 91p² 91d² 91f² 92s² 92p² 92d² 92f² 93s² 93p² 93d² 93f² 94s² 94p² 94d² 94f² 95s² 95p² 95d² 95f² 96s² 96p² 96d² 96f² 97s² 97p² 97d² 97f² 98s² 98p² 98d² 98f² 99s² 99p² 99d² 99f² 100s² 100p² 100d² 100f² 101s² 101p² 101d² 101f² 102s² 102p² 102d² 102f² 103s² 103p² 103d² 103f² 104s² 104p² 104d² 104f² 105s² 105p² 105d² 105f² 106s² 106p² 106d² 106f² 107s² 107p² 107d² 107f² 108s² 108p² 108d² 108f² 109s² 109p² 109d² 109f² 110s² 110p² 110d² 110f² 111s² 111p² 111d² 111f² 112s² 112p² 112d² 112f² 113s² 113p² 113d² 113f² 114s² 114p² 114d² 114f² 115s² 115p² 115d² 115f² 116s² 116p² 116d² 116f² 117s² 117p² 117d² 117f² 118s² 118p² 118d² 118f² 119s² 119p² 119d² 119f² 120s² 120p² 120d² 120f² 121s² 121p² 121d² 121f² 122s² 122p² 122d² 122f² 123s² 123p² 123d² 123f² 124s² 124p² 124d² 124f² 125s² 125p² 125d² 125f² 126s² 126p² 126d² 126f² 127s² 127p² 127d² 127f² 128s² 128p² 128d² 128f² 129s² 129p² 129d² 129f² 130s² 130p² 130d² 130f² 131s² 131p² 131d² 131f² 132s² 132p² 132d² 132f² 133s² 133p² 133d² 133f² 134s² 134p² 134d² 134f² 135s² 135p² 135d² 135f² 136s² 136p² 136d² 136f² 137s² 137p² 137d² 137f² 138s² 138p² 138d² 138f² 139s² 139p² 139d² 139f² 140s² 140p² 140d² 140f² 141s² 141p² 141d² 141f² 142s² 142p² 142d² 142f² 143s² 143p² 143d² 143f² 144s² 144p² 144d² 144f² 145s² 145p² 145d² 145f² 146s² 146p² 146d² 146f² 147s² 147p² 147d² 147f² 148s² 148p² 148d² 148f² 149s² 149p² 149d² 149f² 150s² 150p² 150d² 150f² 151s² 151p² 151d² 151f² 152s² 152p² 152d² 152f² 153s² 153p² 153d² 153f² 154s² 154p² 154d² 154f² 155s² 155p² 155d² 155f² 156s² 156p² 156d² 156f² 157s² 157p² 157d² 157f² 158s² 158p² 158d² 158f² 159s² 159p² 159d² 159f² 160s² 160p² 160d² 160f² 161s² 161p² 161d² 161f² 162s² 162p² 162d² 162f² 163s² 163p² 163d² 163f² 164s² 164p² 164d² 164f² 165s² 165p² 165d² 165f² 166s² 166p² 166d² 166f² 167s² 167p² 167d² 167f² 168s² 168p² 168d² 168f² 169s² 169p² 169d² 169f² 170s² 170p² 170d² 170f² 171s² 171p² 171d² 171f² 172s² 172p² 172d² 172f² 173s² 173p² 173d² 173f² 174s² 174p² 174d² 174f² 175s² 175p² 175d² 175f² 176s² 176p² 176d² 176f² 177s² 177p² 177d² 177f² 178s² 178p² 178d² 178f² 179s² 179p² 179d² 179f² 180s² 180p² 180d² 180f² 181s² 181p² 181d² 181f² 182s² 182p² 182d² 182f² 183s² 183p² 183d² 183f² 184s² 184p² 184d² 184f² 185s² 185p² 185d² 185f² 186s² 186p² 186d² 186f² 187s² 187p² 187d² 187f² 188s² 188p² 188d² 188f² 189s² 189p² 189d² 189f² 190s² 190p² 190d² 190f² 191s² 191p² 191d² 191f² 192s² 192p² 192d² 192f² 193s² 193p² 193d² 193f² 194s² 194p² 194d² 194f² 195s² 195p² 195d² 195f² 196s² 196p² 196d² 196f² 197s² 197p² 197d² 197f² 198s² 198p² 198d² 198f² 199s² 199p² 199d² 199f² 200s² 200p² 200d² 200f² 201s² 201p² 201d² 201f² 202s² 202p² 202d² 202f² 203s² 203p² 203d² 203f² 204s² 204p² 204d² 204f² 205s² 205p² 205d² 205f² 206s² 206p² 206d² 206f² 207s² 207p² 207d² 207f² 208s² 208p² 208d² 208f² 209s² 209p² 209d² 209f² 210s² 210p² 210d² 210f² 211s² 211p² 211d² 211f² 212s² 212p² 212d² 212f² 213s² 213p² 213d² 213f² 214s² 214p² 214d² 214f² 215s² 215p² 215d² 215f² 216s² 216p² 216d² 216f² 217s² 217p² 217d² 217f² 218s² 218p² 218d² 218f² 219s² 219p² 219d² 219f² 220s² 220p² 220d² 220f² 221s² 221p² 221d² 221f² 222s² 222p² 222d² 222f² 223s² 223p² 223d² 223f² 224s² 224p² 224d² 224f

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Table of Contents Answers For Chemquest

1. Understanding the eBook Answers For Chemquest
 - The Rise of Digital Reading Answers For Chemquest
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers For Chemquest
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers For Chemquest
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers For Chemquest
 - Personalized Recommendations
 - Answers For Chemquest User Reviews and Ratings
 - Answers For Chemquest and Bestseller Lists
5. Accessing Answers For Chemquest Free and Paid eBooks
 - Answers For Chemquest Public Domain eBooks

- Answers For Chemquest eBook Subscription Services
- Answers For Chemquest Budget-Friendly Options
- 6. Navigating Answers For Chemquest eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers For Chemquest Compatibility with Devices
 - Answers For Chemquest Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers For Chemquest
 - Highlighting and Note-Taking Answers For Chemquest
 - Interactive Elements Answers For Chemquest
- 8. Staying Engaged with Answers For Chemquest
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers For Chemquest
- 9. Balancing eBooks and Physical Books Answers For Chemquest
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers For Chemquest
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers For Chemquest
 - Setting Reading Goals Answers For Chemquest
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers For Chemquest
 - Fact-Checking eBook Content of Answers For Chemquest
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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