

SCH4U Unit 4 - Quiz – Equilibrium

Player: _____

Period: _____

Q1: (10)

Q2: (10)

Q3: (10)

Q4: (10)

Q5: (10)

Write the formulas for:

$$K_{eq} = \frac{[products]}{[reactants]} \quad ; \quad K_{eq} = \frac{P_{products}}{P_{reactants}}$$

$$K_{eq} = \frac{P_{CO_2}^2 P_{H_2O}}{P_{CH_4}^2 P_{O_2}^2} \quad ; \quad K_{eq} = \frac{P_{CO_2}^2 P_{H_2O}}{(P_{CH_4})^2 (P_{O_2})^2}$$

gpt (aq) + h₂o (l) <=> gpt(s) + h₂g (g) <=> gpt(s) + h₂g (g) <=> gpt(s) + h₂g (g)

h₂o (l) + h₂g (g)

Question 1:

1. Hydrogen sulfide, H₂S, is a slightly poisonous gas. At 100°C, an equilibrium mixture contains 0.020 moles of hydrogen, H₂, 0.010 moles of hydrogen sulfide, and an undetermined amount of sulfur, S₂. The reaction is:



If the value of K is 0.00010 at 100°C, what concentration of S₂ is present at equilibrium?

a. 0.00010

b. 0.00010 moles/litre

c. 0.00010 moles/litre

d. 0.00010 g/l

e. Determine the value of S₂ at equilibrium

f. 0.00010

What other species have chemical formula consisting of 2 atoms each?

g. 0.00010

	2H ₂ (g)	S ₂ (g)	2H ₂ S(g)
Initial (moles)			
Change (moles)			
Final (moles)	0.020	0.010	0

f. 0.00010 moles/litre

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