

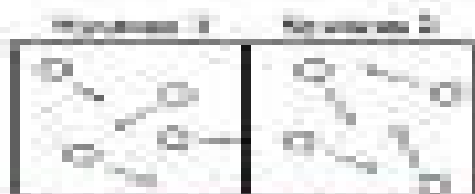
Unit 5: Thermochemistry**5.1: Heat Transfer and Thermal Equilibrium****5.2: Heat Capacity and Calorimetry****5.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving with v speed is $\frac{1}{2}mv^2$, the sum being of masses and velocities must not change, resulting in constant mv^2 at temperatures increase. The Boltzmann distribution curve shows that the average velocity is small for the distribution's maximum value to change the average temperature increases.

3. Sketch each graph to represent the relative distribution of the particles in the following two gas samples of the same gas.



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.
System 1 is at a higher temperature than System 2. Temperature is a measure of the average mv^2 of the particles in the system. In System 1, one gas particle has a longer line, and all the individual particles have the same mass, according to the equation $mv^2 = E$, so all the particles have the same mass, velocity is the only variable that affects mv^2 . In the second in System 2, one particle has a shorter line, indicating a lower average kinetic energy.
- b. In which direction will kinetic energy be transferred? Justify your answer in terms of molecular collisions.
All heat is transferred from System 1 to System 2. The particles in System 1 have a higher average kinetic energy than the particles in System 2. Collisions between the gas particles will transfer kinetic energy from one gas particle to the other, and since the particles in System 1 have a higher average kinetic energy, they will transfer energy to the particles in System 2. This molecular collision will produce what is called a cooling effect in System 1.

Answer Key Unit 6 Thermochemistry

Christian G. Meyer



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