

“PERIODIC TABLE PUZZLE” ANSWER KEY

1. **C** (Cannot be “I” because “I” is Hydrogen, which is a gas, not a metal.)
2. **F** (Group 2 is the Alkaline Earth Metal Family)
3. **A** (“Inactive” means *not reactive*, also known as “inert.” The Noble Gasses in Group 18 are inert because they do not easily react with other elements.)
4. **H** (“Active,” meaning “reactive,” Group 17, the Halogens, are the most reactive nonmetals. Letter “H” occupies the space of the element Fluorine, which is the most active nonmetal on the Periodic Table.)
5. **B** (“Semi-metal” is the same as a “metalloid,” referring to elements that have some of the physical and chemical properties typical of metals, but other properties that are not characteristic of metals.)
6. **D** (“E” is technically a Transition Metal, not an “Inner” Transition Metal)
7. **G** (Oxidation state, or oxidation number, refers to the charge an atom would have if it became an ion. Elements in Group 16, such as “G,” all have 6 valence electrons, meaning they would need 2 more electrons to complete their valence shell. Since electrons have negative charges, gaining 2 more would give the ion a charge of -2.)
8. **E** (Transition Metals like “E” do not follow the standard rules. For instance, many have more than one oxidation number. Also, answer can’t be “B” because B is a semimetal/metalloid, not a metal.)
9. **J** (“B” is in Group 13, meaning it has 3 valence electrons. It would be easier for it donate 3 negative electrons, leaving it with a charge of +3, than it would for it to gain 5, for a charge of -5.)
10. **I** (“I” is in the place of Hydrogen, which could gain 1 electron to fill the first electron shell, or lose 1 electron to become the common positive ion H^+ .)

Answer To Chemistry A Study Of Matter Periodic Puzzle

Victor M. Corman



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