

Chapter 16 Thermal Energy and Heat

Section 16.1 Thermal Energy and Matter

(pages 474–478)

This section defines heat and describes how work, temperature, and thermal energy are related to heat. It also discusses thermal expansion and contraction of materials, and explains uses of a calorimeter.

Reading Strategy (page 474)

Previewing Before you read, preview the figures in this section and add two more questions to the table. As you read, write answers to your questions. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

Thermal Energy and Matter	
Questions About Thermal Energy and Matter	Answers
Which has more thermal energy, a cup of tea or a pitcher of juice?	A pitcher of juice
Why did Rumford conclude that heat is not a form of matter? (Fig. 1)	The brass was hot enough to make water boil only during drilling, so the heat must be related to the motion of the drill.
How is specific heat related to temperature? (Fig. 3)	The lower a material's specific heat, the more its temperature rises when a given amount of energy is absorbed by a given mass.

Work and Heat (page 474)

- Heat is the transfer of thermal energy from one object to another as the result of a difference in _____. Circle the correct answer.
density potential energy **temperature**
- Circle the letter of each sentence that is true about heat.
 - Heat is a fluid that flows between particles of matter.
 - b.** Heat flows spontaneously from hot objects to cold objects.
 - c.** Friction produces heat.

Temperature (page 475)

- What is temperature? Temperature is a measure of how hot or cold an object is compared to a reference point.

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