

Lab 8B: Calculations with a Chemical Structure

Objectives:

1. To derive an empirical formula from a chemical structure without performing stoichiometric calculations.
2. To calculate the number of moles of a solid from a weighing without performing stoichiometric calculations.
3. To determine the molar mass of a solid.
4. To determine the theoretical amount of product produced.
5. To compare the theoretical amount to the actual amount of product formed and calculate the percent yield.

Reaction equation:

Reaction of sodium chloride solution: 25.0 mL

Reaction of calcium chloride solution: 25.0 mL

Reaction of 0.1M NaOH: 5.00 g

Reaction of 0.1M NaOH: 5.00 g

Reaction of 0.1M NaOH: 5.00 g

Reaction of 0.1M NaOH: 5.00 g

Procedure:

Reaction and Calculations:

1. Reaction:

$$\text{moles of sodium chloride} = 0.1 \text{ mol/L} \times 0.025 \text{ L} = 0.0025 \text{ mol}$$

$$\text{moles of calcium chloride} = 0.1 \text{ mol/L} \times 0.025 \text{ L} = 0.0025 \text{ mol}$$

$$\text{moles of sodium hydroxide produced} = 0.0025 \text{ mol} \times 2 = 0.005 \text{ mol}$$

$$\text{moles of product sodium hydroxide} = 0.005 \text{ mol} \times 40 \text{ g/mol} = 0.2 \text{ g}$$

2. Reaction equation: $\text{Na}_2\text{CO}_3 + \text{CaCl}_2 \rightarrow \text{CaCO}_3 + 2\text{NaCl}$

3. Reaction equation:

$$\text{moles of } \text{CaCl}_2 = 0.0025 \text{ mol} \times 111 \text{ g/mol} = 0.2775 \text{ g}$$

$$\text{moles of } \text{Na}_2\text{CO}_3 = 0.0025 \text{ mol} \times 106 \text{ g/mol} = 0.265 \text{ g}$$

$$\text{moles of } \text{CaCl}_2 \text{ and } \text{Na}_2\text{CO}_3 = 0.0025 \text{ mol} \times 111 \text{ g/mol} = 0.2775 \text{ g}$$

Answers For Lab 6b Heath Chemistry

Brendan G. Carr



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