



Name: Chelsey Williams Date: 3/31/22

Student Exploration: Stoichiometry

Vocabulary: Avogadro's number, balanced equation, cancel, coefficient, conversion factor, dimensional analysis, formula mass, molar mass, mole, molecular mass, stoichiometry

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

1. A 250 mL glass of orange juice contains 22 grams of sugar. How much sugar is in a two-liter (2,000 mL) bottle of orange juice? 176 grams
2. It requires two sticks of butter to make a batch of 20 cookies. How much butter will it take to make 150 cookies? 15 sticks of butter

Gizmo Warm-up

Just as a cook follows a recipe to decide how much of each ingredient to add, a chemist uses **stoichiometry** to determine the amounts of substances involved in chemical reactions. The Stoichiometry Gizmo allows you to try your hand at figuring out the amounts of reactants and products that take part in a chemical reaction.

To begin, check that this equation is shown:



1. Look at the **coefficients** (such as the "3" in 3CO) in front of each substance in the equation. The coefficients tell you how many molecules or atoms take part in a chemical reaction. In the spaces below, list the number of each molecule or atom in the equation:

Fe_2O_3 1

CO 3

Fe 2

CO_2 3

2. In a **balanced equation**, the same number of each kind of atom is shown on each side of the equation. Calculate the number of iron (Fe), oxygen (O), and carbon atoms (C).

New equation



New question

How many moles of carbon monoxide (CO) are required to react completely with 1.75 moles of iron (III) oxide (Fe_2O_3)?

Problems solved: 0



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groups of students who are not always well represented in STEM education To influence practice it is important to understand how researchbased practice is made and how it is implemented The volume should be considered as a first step in thinking through what reform in undergraduate science teaching might look like and how we help faculty to implement such reform

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web symbol for candium or whatever you used being careful not to use a symbol that is already in use such as cd or cn for
 your atomic number do not use one that is already in use such as 98 use the average atomic mass that you calculated in lab 2
 see the background information section at the beginning of this lab how many neutrons each do

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 what was the average atomic mass for their candium sample color percent abundance mass of one piece red 0 249 green 0
 358 blue 0 403 2 32 g 2 35 g 2 359 2 219 2 31 g 2 29 g 2 33 g

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web 1 obtain a sample of candium separate the 3 isotopes candium red hots mini candium peanut butter regular and candium peanut large and record the numbers for each type of isotope 2 next determine the mass for each isotope and calculate the total mass in grams 3 calculate the average mass

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