

SCH4U Unit 4 - Quiz – Equilibrium

Name: _____ Period: _____

1201-1202

1203

1204

1205

1206

Equilibrium constants:

$$K_c = \frac{[\text{products}]}{[\text{reactants}]} \quad K_p = \frac{(P)^{\text{products}}}{(P)^{\text{reactants}}}$$

$$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b} \quad K_p = \frac{(P_C)^c (P_D)^d}{(P_A)^a (P_B)^b}$$

gases: $1 \text{ mol} = 22.4 \text{ L}$ $1 \text{ mol} = 22.4 \text{ L}$ $1 \text{ mol} = 22.4 \text{ L}$ $1 \text{ mol} = 22.4 \text{ L}$ $1 \text{ mol} = 22.4 \text{ L}$

Question 1:

1. Hydrogen sulfide, H_2S , is a slightly poisonous gas. At 1000°C, an equilibrium mixture contains 0.020 mol of H_2S , 0.010 mol of hydrogen, H_2 , and 0.010 mol of sulfur, S_2 , in a 1.00 L container. The reaction is:



If the volume of the container is changed to 2.00 L, what concentration of H_2S is present at equilibrium?

- 0.020 mol/L
- 0.010 mol/L
- 0.005 mol/L
- 0.0025 mol/L
- 0.00125 mol/L
- 0.000625 mol/L
- 0.0003125 mol/L
- 0.00015625 mol/L
- 0.000078125 mol/L
- 0.0000390625 mol/L
- 0.00001953125 mol/L
- 0.000009765625 mol/L
- 0.0000048828125 mol/L
- 0.00000244140625 mol/L
- 0.000001220703125 mol/L
- 0.0000006103515625 mol/L
- 0.00000030517578125 mol/L
- 0.000000152587890625 mol/L
- 0.0000000762939453125 mol/L
- 0.00000003814697265625 mol/L
- 0.000000019073486328125 mol/L
- 0.0000000095367431640625 mol/L
- 0.00000000476837158203125 mol/L
- 0.000000002384185791015625 mol/L
- 0.0000000011920928955078125 mol/L
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