

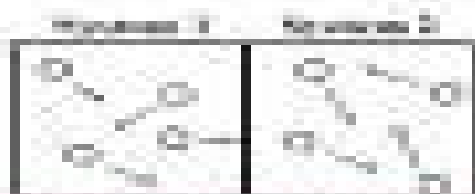
Unit 6: Thermochemistry**6.1: Heat Transfer and Thermal Equilibrium****6.2: Heat Capacity and Calorimetry****6.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving with v is $\frac{1}{2}mv^2$, so the more of particles that move fast enough, the more kinetic energy we will see for a system. Therefore, the Boltzmann-probability curves show that the average velocity we will see for the distribution is increasing (faster to longer) as the temperature increases.

3. Sketch what you expect to represent the relative velocities of the particles in the following two gases, samples of the same gas.



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.
System 1 is at a lower temperature than System 2. Temperature is a measure of the average KE of the particles in the system. In System 1, the particles are moving slower than in System 2, so the kinetic energy is lower. Therefore, the temperature is lower. The particles in System 1 are moving slower than the particles in System 2, so the average KE is lower. The particles in System 2 are moving faster than the particles in System 1, so the average KE is higher.
- b. Are the particles in System 1 moving faster or slower than the particles in System 2? Justify your answer in terms of molecular collisions.
The particles in System 1 are moving slower than the particles in System 2. The particles in System 1 are moving slower than the particles in System 2, so the average KE is lower. The particles in System 2 are moving faster than the particles in System 1, so the average KE is higher. The particles in System 1 are moving slower than the particles in System 2, so the average KE is lower. The particles in System 2 are moving faster than the particles in System 1, so the average KE is higher.

Answer Key Unit 6 Thermochemistry

Michal Rosen-Zvi



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